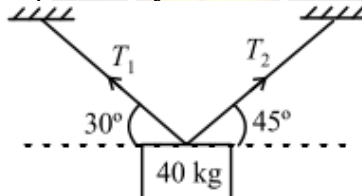
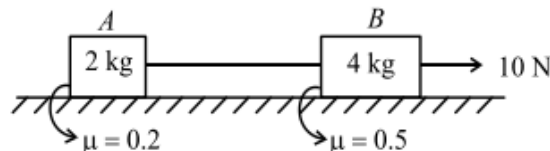




- The displacement of a particle is given as $X = \frac{P}{Q}(1 - e^{-Qt})$, where t is time and P and Q are dimensional constant. The dimension of PQ will be
A) $[M^0LT^{-1}]$ B) $[M^0LT^{-2}]$ C) $[MLT^{-2}]$ D) $[M^0L^{-2}T]$
- The least count of a stopwatch is 0.2s. The time of 100 oscillations is found to be 50s. The maximum error in the measurement is
A) 0.4% B) 0.2% C) 2% D) 4%
- A particle thrown up vertically reaches at half of the maximum height in time $(\sqrt{2} - 1)$ s. The speed of projection of the particle is ($g = 10 \text{ m/s}^2$)
A) $5\sqrt{2} \text{ m/s}$ B) 10 m/s C) 40 m/s D) $10\sqrt{2} \text{ m/s}$
- Virat Kohli can throw a ball to a maximum horizontal distance of 50 m. How much high above the ground can he throw the same ball?
A) 25 m B) 50 m C) 100 m D) 75 m
- If $\vec{A} = 2\hat{i} + 3\hat{j}$ and $\vec{B} = 3\hat{i} + 3\hat{j} + 2\hat{k}$, then the unit vector in the direction of $\vec{A} + \vec{B}$ is
A) $\frac{5\hat{i} + 5\hat{j} + 3\hat{k}}{\sqrt{59}}$ B) $\frac{5\hat{i} + 5\hat{j} - 3\hat{k}}{\sqrt{59}}$ C) $\frac{5\hat{i} + 6\hat{j} - 2\hat{k}}{\sqrt{65}}$ D) $\frac{5\hat{i} + 6\hat{j} + 2\hat{k}}{\sqrt{65}}$
- A body of a mass 40 kg is suspended by two massless strings as shown in the figure, then



- A) $\sqrt{2}T_1 - \sqrt{3}T_2 = 0$ B) $\sqrt{3}T_1 - \sqrt{2}T_2 = 0$ C) $\sqrt{3}T_1 + \sqrt{2}T_2 = 0$ D) $3T_1 - 2T_2 = 0$
- Two blocks A and B are connected with an ideal string are pulled horizontally by a force of 10 N as shown in the figure. The force of friction acting on block A is

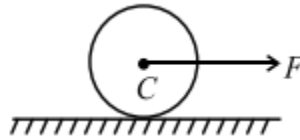


- A) 4 N B) 10 N C) 6 N D) Zero
- The work done by an agent applying a force $\vec{F} = (\hat{i} + 2\hat{j} + 3\hat{k}) \text{ N}$ on a particle in moving it from A(0, 0, 0) m to B(3, 4, 5) m is
A) 26 J B) 20 J C) 13 J D) 39 J
 - Kinetic energy of any moving particle is
A) Scalar, positive B) Scalar, may be negative
C) May be vector, positive D) Vector, may be zero





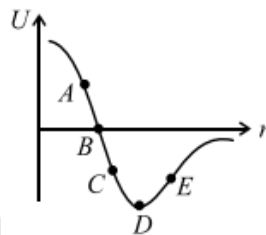
10. A solid sphere placed on a rough horizontal surface, is pulled horizontally by a force F . If it undergoes pure rolling, then the frictional force developed is



- A) $\frac{5}{7}F$ B) $\frac{3}{7}F$ C) $\frac{3F}{5}$ D) $\frac{2}{7}F$

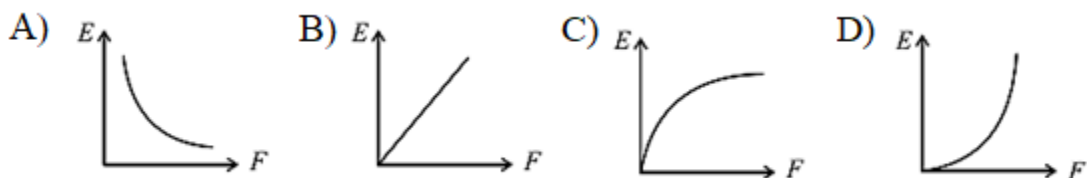
11. The curve between potential energy (U) and distance (r) between atoms of a diatomic molecule

is shown in the figure. If $\vec{F}_A, \vec{F}_B, \vec{F}_C, \vec{F}_D$ and \vec{F}_E are forces of interaction between atoms of molecules corresponding to points shown on curve, then



- A) $|\vec{F}_B| = 0$ B) $|\vec{F}_A| = 0$ C) $\vec{F}_E \cdot \vec{F}_C > 0$ D) $\vec{F}_C \cdot \vec{F}_E < 0$

12. Elongation (E) of a steel wire varies with the elongation force (F) according to the graph (within elastic limit)



13. Pressure inside the two soap bubbles are 1.01 atm and 1.02 atm. The ratio of their free surface area is

- A) 2 : 1 B) 1 : 8 C) 101 : 102 D) 4 : 1

14. If the temperature of the liquid rises, then its coefficient of viscosity

- A) Increases B) Decreases C) Remains same D) Increases and decreases periodically

15. An anisotropic material has coefficient of linear expansion $\alpha, 2\alpha, 2\alpha$ along x, y and z-axis respectively. Then coefficient of cubical expansion is

- A) 2α B) 3α C) 5α D) $\frac{5}{3}\alpha$

16. In the spectrum of a black body at two temperature T and $2T$, let A_1 and A_2 be area under the two curves respectively. The value of $\frac{A_1}{A_2}$ will be

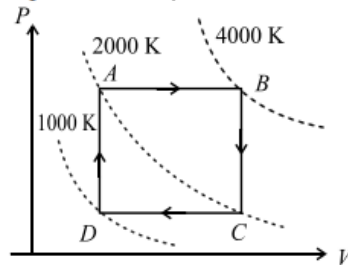
- A) 1 : 16 B) 4 : 1 C) 8 : 1 D) 2 : 1

17. An ideal diatomic gas undergoes a cyclic process as shown in P - V diagram. The dotted curves are





isothermal. The efficiency of the heat engine based on these processes is

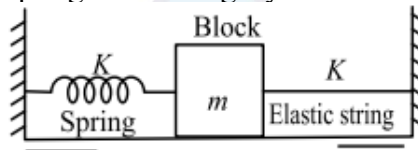


- A) $\frac{2}{13}$ B) $\frac{2}{19}$ C) $\frac{1}{5}$ D) $\frac{1}{7}$

18. The rms speed of an ideal diatomic gas at temperature T is V . When gas dissociates into atoms then its new rms speed becomes double. The temperature at which the gas dissociated into atoms are:

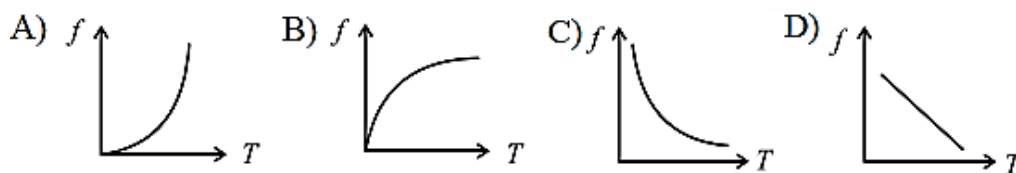
- A) T B) $\sqrt{2}T$ C) $\frac{T}{2}$ D) $2T$

19. The time period of a given spring-mass-string system is

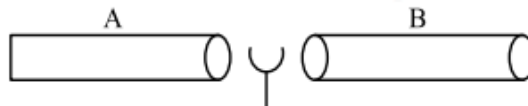


- A) $2\pi\sqrt{\frac{m}{2K}}$ B) $2\pi\sqrt{\frac{m}{K}}$ C) $(\sqrt{2}+1)\pi\sqrt{\frac{m}{2K}}$ D) $(\sqrt{2}+1)\pi\sqrt{\frac{m}{K}}$

20. If graph between fundamental frequency (f) and corresponding tension (T) in a sonometer wire is plotted, then it is best represented by

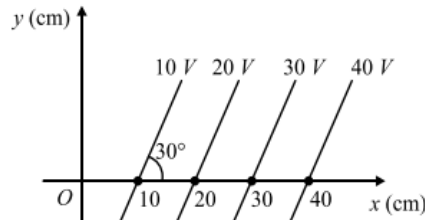


21. A tuning fork is kept between two organ pipes A and B. organ pipe A is closed at one end and is of length 18 cm while organ pipe B is opened at both ends as shown in the figure. If fourth over-tone of pipe B and first overtone of pipe A are in resonance with tuning fork, then the length of open organ pipe B is



- A) 60 cm B) 50 cm C) 40 cm D) 30 cm
22. An infinite number of charges, each of charge $1 \mu\text{C}$ are placed on x-axis with co-ordinates $X(\text{m}) = 1, 2, 4, 8, \dots, \infty$. If a charge 2C is placed at origin, then the net force on 2C charge is
- A) 18000 N B) 24000 N C) 48000 N D) 72000 N
23. The figure shows some of the equipotential surfaces. The magnitude and direction of the electric field given by





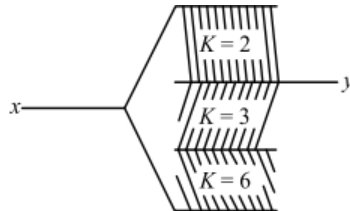
A) 200 V/m, -60° with + x-axis

B) 200 V/m, 120° with + x-axis

C) 100 V/m, 30° with + x-axis

D) 100 V/m, 120° with + x-axis

24. In the given arrangement of parallel plates. Each plate has area A and distance between two consecutive plates is d. The equivalent capacitance of the system between x and y is given as



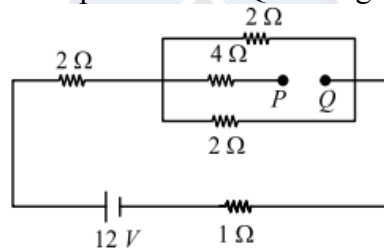
A) $\frac{\epsilon_0 A}{d}$

B) $\frac{2\epsilon_0 A}{d}$

C) $\frac{3\epsilon_0 A}{d}$

D) $\frac{4\epsilon_0 A}{d}$

25. The potential difference between points P and Q in the given circuit is



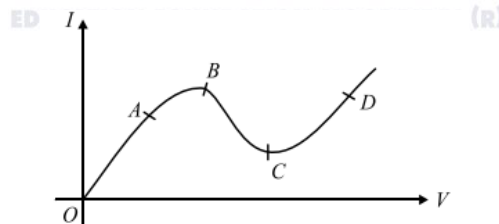
A) 2 V

B) 1 V

C) 3 V

D) 4 V

26. The V-I characteristics of a non-linear device is shown in the figure. The region showing negative resistance is



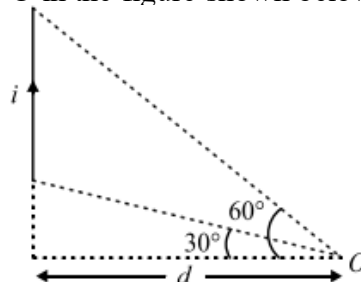
A) OA

B) AB

C) CD

D) BC

27. The magnetic field at point O in the figure shown below is



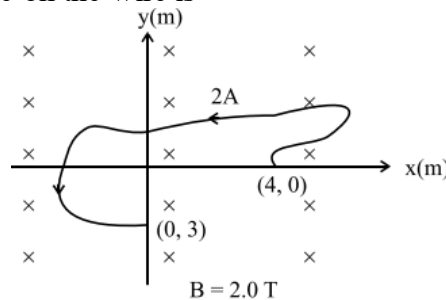


- A) $\frac{\mu_0 i}{4\pi d} [\sqrt{3}-1]$ B) $\frac{\mu_0 i}{8\pi d}$ C) $\frac{\mu_0 i}{8\pi d} (\sqrt{3}-1)$ D) Zero

28. The ratio of magnetic length to geometrical length of a bar magnet is about

- A) 0.6 B) 0.84 C) 0.16 D) 1

29. The end points of a current-carrying wire lie on the x-axis and y-axis as shown in the figure. The magnetic force on the wire is



- A) 10 N B) 20 N C) 30 N D) Zero

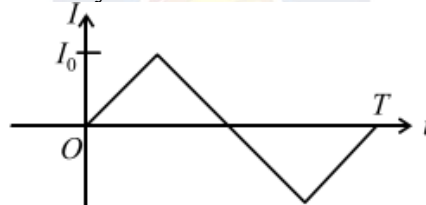
30. Consider the following statements

- (a) Eddy current loss is minimized by using laminated cores
(b) Self-inductance is called inertia of electricity.

Choose the correct statements.

- A) Only a B) Only b C) Both a and b D) Neither a nor b

31. The variation of an alternating current (I) varies with time (t) as shown in the figure. The average value of current for half cycle is



- A) Zero B) $\frac{I_0}{2}$ C) $\frac{2I_0}{\pi}$ D) $\frac{I_0}{\pi}$

32. Which one of the following rays (or wave) has maximum speed in air?

- A) β -rays B) α -rays C) Sound wave D) Heat radiations

33. In compound microscope, intermediate image formed is

- A) Virtual, erect and magnified B) Real, erect and diminished
C) Real, inverted and magnified D) Virtual, erect and diminished

34. A convex mirror is dipped in a liquid whose refractive index is equal to the refractive index of material of mirror. Then its focal length will

- A) Become zero B) Become infinite
C) Remain unchanged D) Become small but not zero

35. In the double-slit experiment, what should be the width of each slit to obtain 8 maxima of the double

slit pattern within the central maxima of single slit pattern with $d = 3 \text{ mm}$?

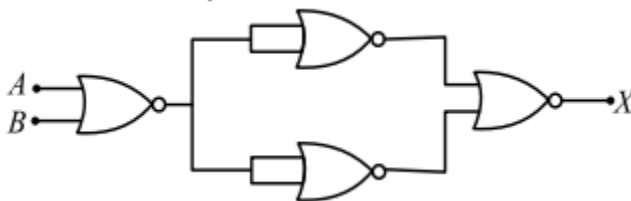
- A) 0.6 mm B) 0.4 mm C) 0.75 mm D) 0.3 mm

36. The work function of a photosensitive surface is 3.2 eV. The wavelength of the incident radiation for which the stopping potential is 3V lies in the





- A) Visible region B) IR region C) UV region D) Radio wave region
37. 1 mg mass is equivalent to
A) 5.66×10^{29} eV B) 4.2×10^{26} eV C) 9×10^{10} eV D) 5.66×10^{26} eV
38. The ratio of longest wavelengths corresponding to Lyman and Balmer series in hydrogen spectrum is
A) $\frac{5}{36}$ B) $\frac{5}{27}$ C) $\frac{5}{9}$ D) $\frac{3}{4}$
39. If a small amount of a trivalent atom is doped with silicon crystal, then
A) Its resistance is increased B) It becomes n-type semiconductor
C) Semiconductor becomes electrically positive D) There will be holes in majority
40. The output (X) of the logic circuit shown in the figure will be



- A) $\overline{A.B}$ B) $\overline{A+B}$ C) $A.B$ D) $A+B$
41. Assertion: In a simple battery circuit, the point of the lowest potential is positive terminal of the battery.
Reason: The current flows towards the point of the higher potential, as it does in such a circuit from the negative to the positive terminal.
A) If both Assertion and Reason are correct and the Reason is a correct explanation of the Assertion.
B) If both Assertion and Reason are correct but Reason is not a correct explanation of the Assertion.
C) If the Assertion is correct but Reason is incorrect.
D) If both the Assertion and Reason are incorrect.
42. Assertion: In electric circuits, wires carrying currents in opposite directions are often twisted together.
Reason: If the wires are not twisted together, the combination of the wires forms a current loop, the magnetic field generated by the loop might affect adjacent circuits or components.
A) If both Assertion and Reason are correct and the Reason is a correct explanation of the Assertion.
B) If both Assertion and Reason are correct but Reason is not a correct explanation of the Assertion.
C) If the Assertion is correct but Reason is incorrect.
D) If both the Assertion and Reason are incorrect.
43. Assertion: Eddy currents are produced in any metallic conductor when magnetic flux is changed around it.
Reason: Electric potential determines the flow of charge.





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NEET (UG)-2025 (CNMT - 3)

QUESTION PAPER

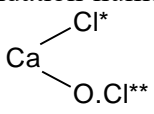
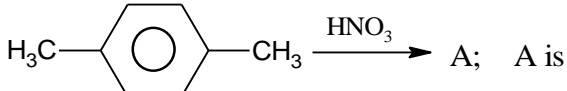
- A) If both Assertion and Reason are correct and the Reason is a correct explanation of the Assertion.
B) If both Assertion and Reason are correct but Reason is not a correct explanation of the Assertion.
C) If the Assertion is correct but Reason is incorrect.
D) If both the Assertion and Reason are incorrect.
44. Assertion: Choke coil is preferred over a resistor to control the current in an AC circuit.
Reason: Power factor of an ideal inductor is zero.
A) If both Assertion and Reason are correct and the Reason is a correct explanation of the Assertion.
B) If both Assertion and Reason are correct but Reason is not a correct explanation of the Assertion.
C) If the Assertion is correct but Reason is incorrect.
D) If both the Assertion and Reason are incorrect.
45. Assertion: The image of an extended object placed perpendicular to the principal axis of a mirror, will be erect if the object is real but the image is virtual.
Reason: The image of an extended object, placed perpendicular to the principal axis of a mirror, will be erect if the object is virtual but the image is real.
A) If both Assertion and Reason are correct and the Reason is a correct explanation of the Assertion.
B) If both Assertion and Reason are correct but Reason is not a correct explanation of the Assertion.
C) If the Assertion is correct but Reason is incorrect.
D) If both the Assertion and Reason are incorrect.

CHEMISTRY

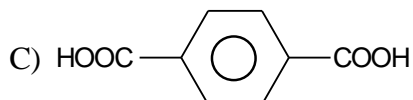
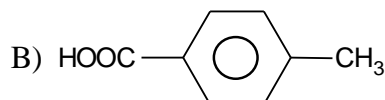
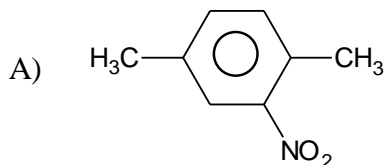
46. 917.65 is rounded off to give answer in four significant figures as
A) 917.6 B) 917.7 C) 917.65 D) 9.176×10^1
47. The number of gram atoms of oxygen in 6.02×10^{24} CO
A) 1 B) 0.5 C) 5 D) 10
48. A near UV photon of 300 nm is absorbed by a gas and then re-emitted as two photons. One photon is red with the wavelength 760 nm. Hence, wavelength of the second photon is
A) 460 nm B) 1060 nm C) 496 nm D) 300 nm
49. For the electrons of oxygen atom, which of the following statements is correct?
A) Z_{eff} for an electron in a 2s-orbital is the same as Z_{eff} for an electron in a 2p-orbital
B) An electron in the 2s-orbital has the same energy as an electron in the 2p-orbital
C) Z_{eff} for an electron in 1s-orbital is the same as Z_{eff} for an electron in a 2s-orbital
D) The two electrons present in the 2s-orbital have spin quantum numbers m_s but of opposite sign
50. Which of the following sequence of arrangement is not in agreement with periodic property written against it?
A) $B < Al < Ga < In$ (atomic radius) B) $C < N < O < F$ (electronegativity)





- C) $\text{Li} < \text{Be} < \text{B} < \text{C}$ (ionisation enthalpy) D) $\text{Be} < \text{Mg} < \text{Ca} < \text{Sr}$ (metallic character)
51. Considering the elements F, Cl, O and N, the correct order of their chemical reactivity in terms of oxidising property is
A) $\text{F} > \text{Cl} > \text{O} > \text{N}$ B) $\text{F} > \text{O} > \text{Cl} > \text{N}$ C) $\text{Cl} > \text{F} > \text{O} > \text{N}$ D) $\text{O} > \text{F} > \text{N} > \text{Cl}$
52. Which of the following molecules or ions is not linear?
A) BeCl_2 B) ICl_2^- C) CS_2 D) ICl_2^+
53. Molecular shapes of SF_4 , CF_4 and XeF_4 are
A) the same, with 2, 0 and 1 lone pair of electrons respectively
B) the same, with 1, 1 and 1 lone pair of electrons respectively
C) different with 0, 1 and 2 lone pairs of electrons respectively
D) different with 1, 0 and 2 lone pairs of electrons respectively
54. Consider the following molecules.
A. Tetracyanomethane B. Carbon dioxide C. Benzene D. 1, 3-butadiene
Ratio of σ and π -bonds are in order
A) $\text{A} = \text{B} < \text{C} < \text{D}$ B) $\text{A} = \text{B} < \text{D} < \text{C}$ C) $\text{A} = \text{B} = \text{D} = \text{C}$ D) $\text{C} < \text{D} < \text{A} < \text{B}$
55. A gas placed in a cylinder fitted with a frictionless piston expands against a constant external pressure of 1 atm from a volume of 5 L to 10 L by absorbing 400 J heat. The change in internal energy is
A) 88.76 J B) -102.22 J C) -106.35 J D) -122.22 J
56. Sign of ΔG is positive at low temperature when
A) ΔH is positive and ΔS is positive B) ΔH is negative and ΔS is negative
C) ΔH is negative and ΔS is positive D) ΔH is negative and ΔS is zero
57. Sulphide ion (S^{2-}) reacts with solid sulphur forming S_2^{2-} and S_3^{2-} with formation constant 12 and 132. Formation constant of S_3^{2-} from sulphur and S_2^{2-} is
A) 12 B) 132 C) 132×12 D) 11
58. Oxidation number of Cl-atoms in CaCl_2 (bleaching powder)

A) zero in each B) -1 in Cl^* and +1 in Cl^{**}
C) +1 in Cl^* and -1 in Cl^{**} D) 1 in each
59. A compound contains atoms of three elements A, B and C. If the oxidation number of A is +2, B is +5 and that of C is -2, the possible formula of the compound is
A) $\text{A}_2(\text{BC}_3)_2$ B) $\text{A}_3(\text{BC}_4)_2$ C) $\text{A}_3(\text{B}_4\text{C})_2$ D) ABC_2
60. The number of primary, secondary, tertiary and quaternary carbons in neo-pentane are respectively,
A) 4,3,2 and 1 B) 5,0,0 and 1 C) 4,0,0 and 1 D) 4,0,1 and 1
61. Consider the following reaction

Choose the correct option



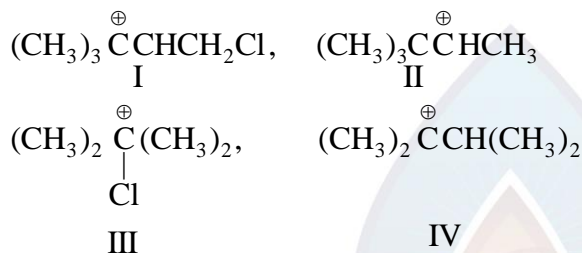


D) None of these

62. Electrophilic addition reactions proceed in two steps. The first step involves the addition of an electrophile. Name the type of intermediate formed in the first step of the following addition reaction. $\text{H}_3\text{C} - \text{HC} = \text{CH}_2 + \text{H}^+ \longrightarrow ?$

A) 2° carbanion B) 1° carbanion C) 2° carbocation D) 1° carbocation

63. Addition of HCl to 3, 3-dimethyl but-1-ene yields two products, one of which has a rearranged carbon skeleton. Which of the following cations are intermediate in that reaction?



A) Both I and II B) Both I and III C) Both II and III D) Both II and IV

64. Basic strength of $\text{CH} \equiv \text{C}^\ominus$ (I); $\text{CH}_2 = \text{CH}^\ominus$ (II) and $\text{CH}_3\text{CH}_2^\ominus$ (III) will be in order

A) I < II < III B) II < III < I C) III < II < I D) III < I < II

65. Acidic nature of CH_3OH (I), CH_3SH (II) and CH_3NH_2 (III) is in order

A) III < II < I B) III < I < II C) I < II < III D) II < I < III

66. Which of the following liquid pairs shows a positive deviation from Raoult's law?

A) water-hydrochloric acid B) benzene-methanol
C) water-nitric acid D) acetone-chloroform

67. A solution of urea (mol mass 56 g mol^{-1}) boils at 100.18°C at the atmospheric pressure. If K_f and K_b for water are 1.86 and $0.512 \text{ K kg mol}^{-1}$ respectively, the above solution will freeze at

A) -6.54°C B) -0.654°C C) 6.54°C D) 0.654°C

68. The conductance of 0.1 M HCl solution is greater than that of 0.1 M NaCl . This is because

A) HCl is more ionized than NaCl
B) HCl is an acid whereas NaCl solution is neutral
C) H^+ ions have greater mobility than Na^+ ions
D) Interionic forces in HCl are weaker than those in NaCl.

69. The emf of the cell $\text{Ti} | \text{Ti}^+ (0.001 \text{ M}) || \text{Cu}^{2+} (0.01 \text{ M}) | \text{Cu}$ is 0.83 V . The emf of this cell could be increased by

A) increasing the concentration of Ti^+ ions B) increasing the concentration of Cu^{2+} ions
C) increasing the concentration of both D) none of the above





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NEET (UG)-2025 (CNMT - 3)

QUESTION PAPER

70. In which of the following cases, does the reaction go farthest to completion?
A) $k = 10^2$ B) $k = 10^{-2}$ C) $k = 10$ D) $k = 1$
71. Which of the following represents the expression for $3/4$ th life of a first order reaction
A) $\frac{k}{2.303} \log 4/3$ B) $\frac{2.303}{k} \log 3/4$ C) $\frac{2.303}{k} \log 4$ D) $\frac{2.303}{k} \log 3$
72. In the P^{3-} , S^{2-} and Cl^- ions the increasing order of size is
A) $Cl^- < S^{2-} < P^{3-}$ B) $P^{3-} < S^{2-} < Cl^-$ C) $S^{2-} < Cl^- < P^{3-}$ D) $S^{2-} < P^{3-} < Cl^-$
73. Which of the following compounds is the strongest reducing agent?
A) H_2O B) H_2S C) H_2Se D) H_2Te
74. The electronegativity follows the order
A) $F > O > Cl > Br$ B) $F > Cl > Br > O$ C) $O > F > Cl > Br$ D) $Cl > F > O > Br$
75. Which of the following transition metals has the highest melting point?
A) Cr B) Mo C) W D) Hg
76. When acidified solution of $K_2Cr_2O_7$ is shaken with aqueous solution of $FeSO_4$, then
A) $Cr_2O_7^{2-}$ ion is reduced to Cr^{3+} ions B) $Cr_2O_7^{2-}$ ion is converted to CrO_4^{2-} ions
C) $Cr_2O_7^{2-}$ ion is oxidised to Cr D) $Cr_2O_7^{2-}$ ion is oxidised to CrO_3
77. If the lanthanoid element with xf electrons has a pink colour, then the lanthanoid with $(14 - x)f$ electrons will have the colour as
A) Blue B) Red C) Green D) Pink
78. The shape of the complex $[Ni(CO)_4]$ is
A) Square planar B) Tetrahedral C) Octahedral D) None of these
79. The complex ions $[Fe(CN)_6]^{3-}$ and $[Fe(CN)_6]^{4-}$
A) Are both octahedral and paramagnetic
B) Are both octahedral and diamagnetic
C) Have same structure but opposite magnetic character
D) Have different structures but same magnetic character.
80. Brown ring test for nitrates is due to the formation of
A) $[Fe(H_2O)_5NO]^{2+}$ B) $[Fe(H_2O)_6]^{2+}$ C) $[Fe(H_2O)(NO)_5]^{2-}$ D) $[Fe(CN)_5NO]^{2+}$
81. The IUPAC name of a secondary optically active alkyl halide having molecular formula, $C_5H_{11}Br$ is
A) 1-bromopentane B) 3-bromopentane C) 2-bromopentane D) 1-bromo-2-methylbutane
82. When silver propanoate is treated with I_2 , in CCl_4 , the major product formed is
A) Iodoethane B) n-Propyl propanoate C) Ethyl propanoate D) 1-Iodoethane
83. Identify 'Z' in the following sequence of reactions
$$C_3H_7I \xrightarrow{KOH alc.} X \xrightarrow[\Delta]{NBS, hv} Y \xrightarrow{KCN aq.} Z$$

A) $(CH_3)_2CH - CN$ B) $Br - CH = CH - CN$
C) $CH_2 = CH - CH_2CN$ D) $CH_2 = CH - CHBr - CN$
84. For which one of the following pairs of compounds can the members be distinguished by means of iodoform test?
A) CH_3OH and $CH_3CH_2CH_2OH$ B) CH_3CH_2OH and $CH_3CHOHCH_3$



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- C) $\text{CH}_3\text{CH}_2\text{CH}_2\text{OH}$ and $\text{CH}_3\text{CHOHCH}_3$ D) $\text{C}_6\text{H}_5\text{CH}_2\text{OH}$ and CH_3OH
85. Phenol is less acidic than
A) ethanol B) methanol
C) o-nitrophenol D) p-methylphenol.
86. The products (A) and (B) in the following reactions respectively are
$$\text{PhCH}=\text{CH}_2 \xrightarrow[\text{(ii) NaNH}_2/\text{liq.NH}_3]{\text{(i) Br}_2/\text{CCl}_4} \text{(A)} \xrightarrow[\text{HgSO}_4]{\text{Dil.H}_2\text{SO}_4} \text{(B)}$$

A) PhCH_2CHO , $\text{PhCH}_2\text{CH}_2\text{OH}$ B) $\text{PhC} \equiv \text{CNa}$, PhCOCH_3
C) $\text{PhC} \equiv \text{CH}$, PhCH_2CHO D) PhCHCHCH_3 , PhCOCHO
87. The correct order of reactivity of >C=O group in the following compounds is:
A) $\text{CH}_3\text{CHO} > (\text{CH}_3)_2\text{CO} > \text{C}_2\text{H}_5\text{COCH}_3$ B) $\text{C}_2\text{H}_5\text{COCH}_3 > (\text{CH}_3)_2\text{CO} > \text{CH}_3\text{CHO}$
C) $(\text{CH}_3)_2\text{CO} > \text{CH}_3\text{CHO} > \text{C}_2\text{H}_5\text{COCH}_3$ D) $(\text{CH}_3)_2\text{CO} > \text{C}_2\text{H}_5\text{COCH}_3 > \text{CH}_3\text{CHO}$
88. What is the end product in the following sequence of reactions?
$$\text{C}_6\text{H}_6 \xrightarrow[333\text{ K}]{\text{HNO}_3-\text{H}_2\text{SO}_4} \text{A} \xrightarrow{\text{Sn/HCl}} \text{B}$$

A) Nitrobenzene B) Aniline C) Benzonitrile D) Phenyl carbamate
89. Acetamide and ethylamine can be distinguished by reacting with
A) Aqueous HCl and heat B) Aqueous NaOH and heat
C) Acidified KMnO_4 D) Bromine water.
90. Which of the following is not a pyrimidine?
A) Uracil B) Guanine C) Cytosine D) Thymine

BOTANY

91. House fly belongs to the order
A) Diptera B) Carnivora C) Primata D) Insecta
92. Study the four statements (A-D) given below and select the two correct ones out of them.
A. Definition of biological species was given by Ernst Mayr
B. Photoperiod does not affect reproduction in plants.
C. Binomial nomenclature system was given by R.H. Whittaker
D. In unicellular organism, reproduction is synonymous with growth
The correct statements are
A) C and D B) A and B C) A and D D) B and C
93. Specialized cells called heterocyst are present in
A) Dinoflagellates B) Chrysophytes C) Archaeobacteria D) Cyanobacteria
94. Select the incorrect statement.
A) Bacterial cell wall is made up of peptidoglycan
B) Pili and Fimbriae are mainly involved in motility of bacterial cell
C) Cyanobacteria lack of flagellated cells



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D) Mycoplasma is a wall less microorganism

95. Choose the correct statements about Liverworts.

- I. In liverworts, sexual reproduction occurs by fusion antherozoid and egg, which are produced in antheridium and archegonium respectively.
- II. Both male and female sex organs may be present on same thallus or different thalli.
- III. Zygote gives rise to sporophyte, which is differentiated into foot, seta and capsule.
- IV. Some cells of capsule undergo meiosis and give rise to haploid spores.

A) I, II and III B) II, III and IV C) I, III and IV D) I, II, III and IV

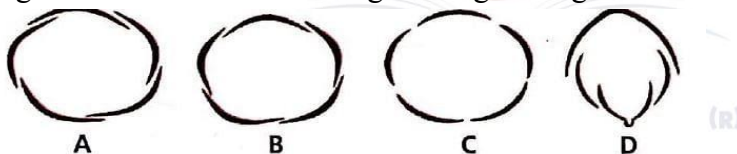
96. Match the following columns.

	Column I		Column II
A.	Red algae	1.	<i>Marchantia</i>
B.	Liverwort	2.	<i>Pinus</i>
C.	Fern	3.	<i>Polysiphonia</i>
D.	Gymnosperm	4.	<i>Adiantum</i>

Codes

	A	B	C	D
A)	1	2	4	3
B)	2	4	3	1
C)	2	3	1	4
D)	3	1	4	2

97. Arrange in correct order according to the given figure.



- A) A-Imbricate, B-Valvate, C-Twisted, D-Vexillary
- B) A-Twisted, B-Imbricate, C-valvate, D-vexillary
- C) A-Valvate, B-Imbricate, C-Twisted, D-Vexillary
- D) A-Vexillary, B-Twisted, C-Imbricate, D-Valvate

98. In which plant the leaf base is swollen?

- A) Monocots B) Leguminous
- C) All dicots D) All angiosperms

99. Match the types of vascular bundles with their description.

	List I		List II
A)	Open vascular bundles	1)	Found in dicotyledonous stem, contain cambium between xylem and phloem, and have the ability to form secondary tissue





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B)	Closed vascular bundles	2)	Do not have cambium, hence do not form secondary tissue, commonly found in monocotyledons
C)	Radial vascular bundles	3)	Xylem and phloem are arranged in different radii, typically found in roots
D)	Conjoint vascular bundles	4)	Xylem and phloem are situated at the same radius, with phloem usually located on the outer side of xylem, common in leaves and stem

	A	B	C	D
A)	1	2	3	4
B)	2	1	3	4
C)	4	3	2	1
D)	1	4	2	3

100. Assertion: Parenchyma cells are found in almost all the parts of the plant and play a role in storage, photosynthesis etc in plants.

Reason: Parenchyma cells have thick cell walls that provide structural support to plants.

- A) Both the assertion and the reason are true and the reason is a correct explanation of the assertion
B) Both the assertion and reason are true but the reason is not a correct explanation of the assertion
C) The assertion is true but the reason is false
D) The assertion is false but the reason is true
101. Find out the incorrect statement from the following.
A) The mitochondria are the power house of the cell because they generate ATP through cellular respiration.
B) Ribosomes are responsible for the synthesis of lipids in the cell.
C) The cell membrane is composed mainly of phospholipid bilayer, which regulates the entry and exit of substances.
D) Chloroplast in plant cell contains chlorophyll, which capture light energy for photosynthesis.
102. The structure inside the chloroplasts that contain stacks of membrane bound sacs is called _____. these stacks are known as _____.
A) Thylakoid, grana
B) Stroma, thylakoid
C) Grana, stroma
D) Ribosomes and thylakoid
103. Statement I: Mitosis results in the production of four genetically identical cells.
Statement II: Meiosis results in the production of four genetically identical cells.
A) Both statement I and II are correct.
B) Only statement I is correct.





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- C) Only statement II is correct.
D) Both the statement I and II are incorrect.
104. Which one of the following statements is true regarding the Go phase?
A) The Go phase is the stage where the cells are actively dividing.
B) Cells in the Go phase are metabolically inactive and do not divide.
C) Go phase can be permanent or reversible, depending on the cell type and conditions.
D) In the Go phase, cells are actively replicating their DNA.
105. Match the following with Column I and column II and choose the correct option.

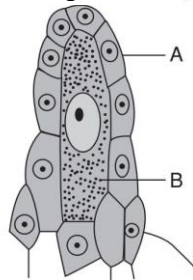
	Column I		Column II
A.	Complex III	i	Succinate dehydrogenase
B.	Complex I	ii	NADH Dehydrogenase
C.	Complex IV	iii	Cytochrome bc1
D.	Complex II	iv	Cytochrome c oxidase

- A) A-iv B-ii C-i D-iii B) A-ii B-i C-iv D-iii
C) A-iii B-ii C-iv D-I D) A-iv b-iii c-ii D-i
106. What is the correct ratio of ATP utilization in steps of the Calvin cycle?
A) Reduction: Regeneration → 1:1 B) Reduction: Regeneration → 2:1
C) Reduction: Regeneration → 2:2 D) Reduction: Regeneration → 1:2
107. ____ A ____ performed an elegant experiment with an ____ B ____ plant to show that in bright sun light, small bubble, were formed around the ____ C ____ parts, while in dark they do not.
- | | | |
|---------------|-------------|---------------|
| A | B | C |
| A) Ingenhousz | Aquatic | Green |
| B) Ingenhousz | Terrestrial | Green |
| C) Priestly | Aquatic | Green |
| D) Priestly | Aquatic | Non-green (R) |
108. Which of the following relations shows substrate level phosphorylation?
A) Citric acid → α-ketoglutaric acid B) Malic acid → oxaloacetic acid
C) α-ketoglutaric acid → Succinyl-CoA D) Succinyl-CoA → Succinic acid
109. Assertion: 2, 4 D is selective weedicide
Reason: 2, 4 D is widely used to kill dicotyledonous weeds and doesn't affect mature monocotyledonous plant
A) Both the assertion and the reason are true and the reason is a correct explanation of the assertion.
B) Both the assertion and reason are true but the reason is not a correct explanation of the assertion.
C) Assertion is true but the reason is false.
D) Both the assertion and reason are false.
110. Match the following
(A) IAA (i) Herring sperm DNA



- (B) ABA (ii) Bolting
(C) Ethylene (iii) Stomatal closure
(D) GA (iv) Weed-free lawns
(E) Cytokinins (v) Ripening of fruits
A) (A) – (iv), (B) – (iii), (C) – (v), (D) – (ii), (E) – (i)
B) (A) – (v), (B) – (iii), (C) – (iv), (D) – (ii), (E) – (i)
C) (A) – (iv), (B) – (i), (C) – (v), (D) – (iii), (E) – (ii)
D) (A) – (v), (B) – (iii), (C) – (ii), (D) – (i), (E) – (iv)

111. Identify the ploidy of the following parts of flowering plants. Ovary, Anther, Egg, Pollen, Male gamete and Zygote
A) $2n, 2n, n, n, 2n$ B) $2n, 3n, n, n, 2n, 2n$ C) $2n, n, n, n, n, 2n$ D) $2n, 2n, n, 2n, n, 2n$
112. How many microsporangia are there in each lobe of anther?
A) One microsporangia B) Two microsporangia
C) Three microsporangia D) Four microsporangia
113. Match the following
- | Column – I | Column – II |
|-----------------------|--|
| A. Vegetative cell | 1. Sporopollenin |
| B. Generative cell | 2. Spindle-shaped cell |
| C. Exine | 3. Large sized and has abundant food reserve |
| D. Intine | 4. Cellulose and pectin |
| A) A:1, B:2, C:3, D:4 | B) A:4, B:3, C:2, D:1 |
| C) A:3, B:2, C:1, D:4 | D) A:2, B:1, C:4, D:3 |
114. Identify the parts A and B in this figure.



- A) A: Megaspore tetrad, B: Nucleus B) A: Central cell, B: Megaspore dyad
C) A: Nucellus, B: Megaspore mother cell D) A: Nucellus, B: Central cell
115. Taylor and his colleagues used which plant to detect the distribution of newly synthesized DNA?
A) *Pisum sativum* B) *Mirabilis jalapa* C) *Vicia faba* D) *Triticum aestivum*
116. β -galactosidase is synthesized by *E. coli* to catalyse hydrolysis of _____ into _____ and glucose.
A) Galactose, lactose B) Galactose, glucose



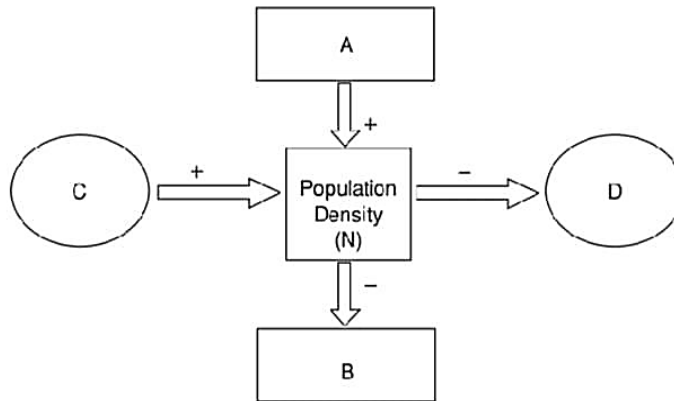
- C) Lactose, galactose D) Maltose, galactose
117. Assertion: Both the strands of DNA are not copied during the process of transcription.
Reason: The two molecules of RNA complementary to each other form the double stranded RNA and this would prevent RNA from being translated into protein.
- A) Both the assertion and the reason are true and the reason is a correct explanation of the assertion.
B) Both the assertion and reason are true but the reason is not a correct explanation of the assertion.
C) Assertion is true but the reason is false.
D) Both the assertion and reason are false.
118. The coding sequences of DNA are known as _____ and the intervening sequences are known as _____ respectively.
- A) Exon, intron B) Intron, exon
C) Cistron, exon D) Exon, cistron
119. BOD can be defined as
- A) Amount of CO₂ consumed if all the inorganic matter in one litre of water is oxidized by bacteria. B) Amount of O₂ consumed if all the organic matter in one decilitre of water is reduced by bacteria. C) Amount of O₂ consumed if all the organic matter in one litre of water is oxidized by bacteria.
D) Amount of O₂ consumed if all the inorganic matter in one litre of water is oxidized by bacteria.
120. Which one of the following is a wrong matching of a microbe and its industrial product, while the remaining three are correct?
- A) Yeast – statins B) *Acetobacter aceti* – acetic acid
C) *Clostridium butylicum* – lactic acid D) *Aspergillus niger* – citric acid
121. Match the following list of microbes and their importance:
- | | | | |
|-----|------------------------------------|-------|---|
| (a) | <i>Saccharomyces cerevisiae</i> | (i) | Production of immunosuppressive agents |
| (b) | <i>Monascus purpureus</i> | (ii) | ripening of Swiss cheese |
| (c) | <i>Trichoderma polysporum</i> | (iii) | Commercial production of ethanol |
| (d) | <i>Propionibacterium sharmanii</i> | (iv) | Production of blood-cholesterol lowering agents |
- (a) (b) (c) (d)
A) (iv) (i) (iii) (ii)
B) (iii) (iv) (i) (ii)
C) (ii) (i) (iv) (iii)
D) (i) (iii) (ii) (iv)
122. Which of the following is true about predators?
- A) They keep the prey population under control.
B) They may help in maintaining species diversity in a community by reducing the intensity of competition among competing prey species.
C) They are prudent in nature.





D) All the above

123. Identify A, B, C and D in this figure.



A) A: Immigration (I), B: Emigration (E), C: Natality (B), D: Mortality (D)

B) A: Natality (B), B: immigration (I), C: Mortality (D), D: Emigration (E)

C) A: Mortality (D), B: Emigration (E), C: Natality (B), D: Immigration (I)

D) A: Mortality (D), B: Natality (B), C: Emigration (E), D: Immigration (I)

124. Connell's elegant field experiments showed that on rocky sea coasts of Scotland, the larger and competitively superior barnacle _____ dominates the intertidal area and exclude smaller barnacle _____ from that zone

A) Chathamalus, Balanus

B) Chathamalus, Chathamalus

C) Balanus, Balanus

D) Balanus, Chathamalus

125. Which of the following are limitations of ecological pyramids?

A) It does not take into account the same species belonging to two or more trophic levels.

B) It assumes a simple food chain that almost does not exist in nature.

C) Saprophytes are not given at any place.

D) All the above

126. Find out the correct statement:

A) Trophic level represents a functional level, not a species as such.

B) A given species never occupies more than one trophic level in the same ecosystem at the same time.

C) In most of the ecosystems, producers are less in number and biomass than the herbivores.

D) Pyramid of energy can never be upright.

127. Match the columns:

Column-I

Column-II

1. Fragmentation

A. Break down of detritus into smaller particles

2. Leaching

B. Precipitation of water-soluble nutrients as unavailable salt





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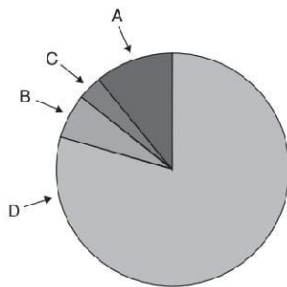
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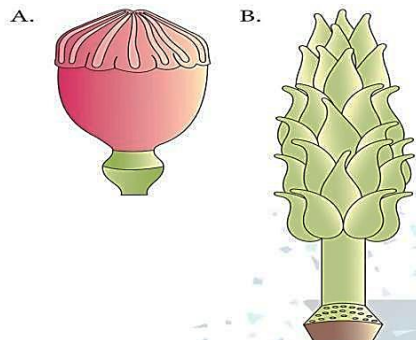
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3. Catabolism C. Degradation of detritus by bacterial and fungal enzymes
- A) 1-A, 2-B, 3-C B) 1-C, 2-B, 3-A C) 1-B, 2-C, 3-A D) 1-C, 2-A, 3-B
128. In which of the following both pairs have the correct combination?
- A) In situ conservation: National park
Ex situ conservation: Botanical Garden
- B) In situ conservation: Cryopreservation
Ex situ conservation: Wildlife Sanctuary
- C) In situ conservation: Seed Bank
Ex situ conservation: National park
- D) In situ conservation: Tissue culture
Ex situ conservation: Sacred groves
129. The species confined to a particular region and not found elsewhere is termed
- A) Alien B) Endemic C) Rare D) Keystone
130. Given below is the representation of the extent of global diversity of invertebrates. What groups the four portions (A to D) represent respectively?



- A) A: Insects, B: Crustaceans, C: Other animal groups, D: Molluscs
- B) A: Crustaceans, B: Insects, C: Molluscs, D: Other animal groups
- C) A: Molluscs, B: Other animal groups, C: Crustaceans, D: Insects
- D) A: Insects, B: Molluscs, C: Crustaceans, D: Other animal groups
131. Casparian strips are made up of ____ and found as depositions in the ____.
- A) Cutin; epidermis B) Lignin; hypodermis
- C) Pectin; exodermis D) Suberin; endodermis
132. Observe the diagrams and identify the correct option that represents the condition and corresponding examples.





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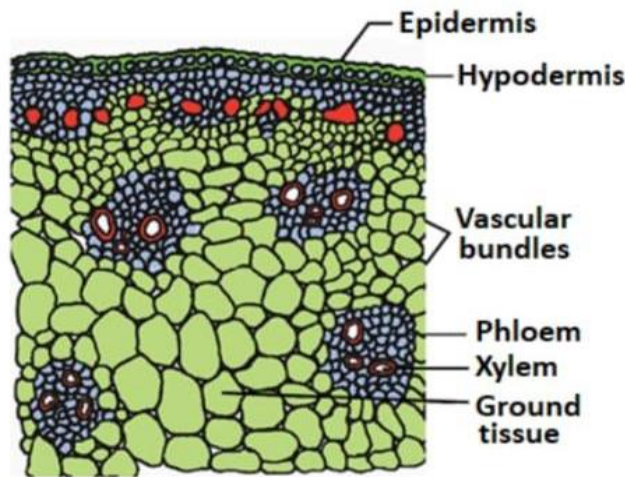
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- A) A – Syncarpous – *Papaver* and B – Apocarpous – *Michelia*
B) A – Apocarpous – *Michelia* and B – Syncarpous – *Papaver*
C) A – Hypogynous – Sunflower and B – Epigynous – China rose
D) A – Perigynous – *Datura* and B – Hypogynous – Peach
133. The floral formula is $\oplus \overset{\square}{O} K_{2+2} C_4 A_{2+4} \underline{G}_{(2)}$ for family:
A) Brassicaceae
B) Solanaceae
C) Malvaceae
D) Anacardiaceae
134. The given transverse section is of



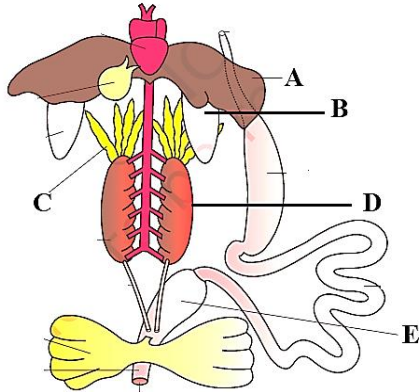
- A) Monocot root
B) Monocot stem
C) Dicot stem
D) Dicot root
135. Specialized epidermal cells surrounding the guard cells are called?
A) bulliform cells
B) lenticels
C) subsidiary cells
D) complementary cells

ZOOLOGY

136. Assertion: Sponges exhibit “cellular level” of body organisation.
Reason: There is some division of labour(activities) occur among the cells
A) Both A and R are correct but R is not the correct explanation of A
B) A is correct but R is not correct
C) R is correct but A is not correct
D) Both A and R are correct and R is the correct explanation of A
137. The mitotic division starts as the zygote moves through the isthmus of the oviduct is
A) Cleavage
B) Morula
C) Trophoblast
D) Implantation
138. The presence of chromogenic substrate gives _____ colonies if the plasmid in a bacterium does not have an insert.
A) No colour
B) Blue colour
C) No colonies
D) White colour



139. The figure given here shows diagrammatic representation of internal organs of frogs. Identify A to E and select the correct option.

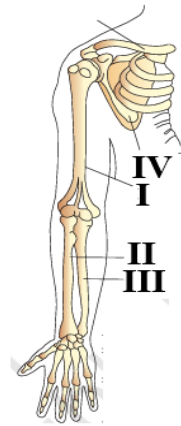


- | A | B | C | D | E |
|-----------------|----------|------------|--------------|--------|
| A) Gall bladder | Liver | Testis | Rectum | Cloaca |
| B) Liver | Lung | Fat Bodies | Kidney | Rectum |
| C) Pancreas | Lung | Testis | Liver | Cloaca |
| D) Liver | Pancreas | Fat Bodies | Gall Bladder | Rectum |
140. Choose the correct pair in reference to the function.
- | | | |
|-------------|---|--------------------------------------|
| A) Insulin | - | Enzyme |
| B) Trypsin | - | Hormone |
| C) GLUT – 4 | - | Enables glucose transport into cells |
| D) Antibody | - | Amino acid |
141. Partial pressure (in mmHg) of oxygen and carbon dioxide at different parts involved in diffusion in comparison to these in atmosphere.

Respiration	Atmospheric air	Alveoli	Blood deoxygenated	Blood oxygenated	Tissue
O ₂	A	104	40	D	40
CO ₂	0.3	B	C	40	45

- | | A | B | C | D |
|----|-----|----|----|----|
| A) | 140 | 40 | 45 | 45 |
| B) | 159 | 45 | 40 | 95 |
| C) | 159 | 40 | 45 | 95 |
| D) | 104 | 45 | 40 | 40 |
142. Globulin is primarily involved in
- | | |
|----------------------|-----------------------|
| A) Allergic reaction | B) Clotting mechanism |
| C) Defense mechanism | D) Osmotic balance |
143. The volume of blood pumped out by each ventricle per minute is called
- | | |
|-------------------|-------------------|
| A) Stroke volume | B) Pace maker |
| C) Cardiac output | D) Joint diastole |

144. Identifying the incorrect statements and choose the correct option accordingly.
- (i) Interstitial fluid (tissue fluid) and lymph have almost similar composition.
 - (ii) Exchange of the nutrients and gases, etc. between blood and cells always occurs through tissue fluid.
 - (iii) Interstitial fluid has the same mineral distribution as that of the plasma.
 - (iv) Lymph can be defined as the blood minus RBC but has specialised lymphocytes.
- A) (i) and (ii) B) (ii) and (iii)
C) (i) and (iv) D) None of the above
145. Statement I : Kidneys are reddish brown and bean – shaped structures.
Statement II : Each kidney of an adult human measures 10 – 12 cm in length, 5 – 7 cm in width, 2 – 3 cm thickness and average weight is 120 – 170 g.
- A) Both Statement I and Statement II are incorrect
B) Statement I is correct but statement II is incorrect
C) Statement II is correct but statement I is incorrect
D) Both statement I and statement II are correct
146. The given diagram represents the bones of the human arm.
-
- The diagram shows the bones of the human arm, including the humerus, radius, ulna, and the bones of the hand and wrist. The bones are labeled with Roman numerals from I to XXIV.



	I	II	III	IV
A)	Clavicle	Ulna	Radius	Humerus
B)	Humerus	Radius	Ulna	Scapula
C)	Scapula	Radius	Ulna	Clavicle
D)	Humerus	Ulna	Radius	Scapula

147. The cerebrum is longitudinally divided into 2 halves that are connected by the
A) Limbic system B) Corpora quadrigemina
C) Corpus callosum D) Cerebral aqueduct
148. Which of the following statements are presented correctly?
i. Alfred Wallace, a naturalist – Malay Archipelago
ii. Charles Darwin – HMS Beagle – Galapagos island
A) Only i is correct B) Only ii is correct
C) Both are correct D) Both are incorrect
149. Column I list the parts of the human brain and column II lists the functions. Match the two columns and identify the correct option.

Column – I		Column - II	
A.	Cerebrum	I	Body temperature
B.	Cerebellum	II	Tracts of nerve fibres
C.	Hypothalamus	III	Controls gastric secretions
D.	Medulla oblongata	IV	Convolutcd surface



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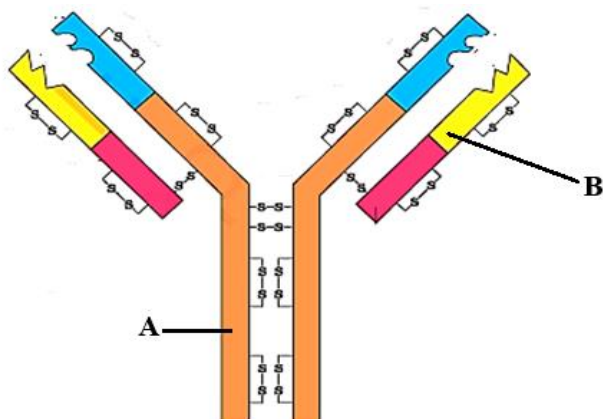
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QUESTION PAPER

- A) A – II, B – IV, C – I, D – III B) A – IV, B – II, C – III, D – I
C) A – I, B – IV, C – III, D – II D) A – III, B – IV, C – I, D – II
150. Which of the following statements is / are correct?
(i) Endocrine cells present in different parts of the gastro – intestinal tract secrete four major peptide hormones, namely gastrin, secretin, cholecystokinin (CCK) and gastric inhibitory peptide (GIP).
(ii) Gastrin acts on the gastric glands and stimulates the secretion of hydrochloric acid and pepsinogen.
(iii) Secretin acts on the exocrine part of the pancreas and stimulates secretion of water and bicarbonate ions.
(iv) GIP inhibits gastric secretion and motility.
(v) Hormones produce their effects on target tissues by binding to specific proteins called hormone receptors located in the target tissues only
A) only (i), (ii) and (v) B) only (ii) and (iv)
C) only (v) D) all of them
151. In the later phase of the pregnancy, the hormone which is also secreted in the ovary is
A) Human placental lactogen B) Oxytocin
C) Progesterone D) Relaxin
152. The corpus luteum secretes large amount of progesterone which is essential for maintenance of
A) Endometrium B) Myometrium
C) Menstrual flow D) Estrogen
153. Which of the following statements is correct regarding the menstrual cycle?
A) LH induces rupturing of graafian follicle
B) Proliferative phase is characterised by the increased production of progesterone.
C) Corpus luteum secretes a large amount of estrogen.
D) Both LH and FSH attain a peak level in the secretory phase.
154. Identify A & B in the following figure.



- A) A – heavy chain, B – light chain B) A – light chain, B – heavy chain
C) A – light chain, B – light chain D) A – heavy chain, B – heavy chain



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155. Guess an odd pair from the following.
- A) Pseudocoelomate - *Ascidia* B) Metamerism - *Nereis*
C) Acoelomate - *Taenia* D) Coelomate - *Laccifer*
156. Choose the correct statement.
- i. According to the WHO, reproductive health is total well-being in the physical, social, emotional and behavioural aspects of reproduction.
ii. According to the WHO, reproductive health is total well-being in the physical, social and emotional aspects of reproduction.
iii. A reproductively healthy society has people with physically and functionally normal reproductive organs
iv. A reproductively healthy society has abnormal sex – related emotional and behavioural interaction.
- A) i, ii and iii B) ii and iv C) i and iii D) i only
157. Statement -A: Reproductive and child healthcare programme is for reproduction related areas.
Statement -B: It deals with creating awareness among various financial related aspects.
- A) Statement – A is correct. B) Statement -B is correct
C) Both statements are correct D) Both statements are incorrect.
158. Morgan and his group also found that even when genes were grouped on the same chromosome, some genes were very tightly linked showed _____ while others were loosely linked showed _____
- A) Higher recombination, lower recombination
B) Lower recombination, lower recombination
C) Very low recombination, higher recombination
D) High recombination, higher recombination
159. Assertion(A): Only boy child could be born with the substitution of glutamic acid by valine on 6th position of β - chain of haemoglobin.
Reason (B): This gene for the above mutation occurs in Y – chromosome
- A) If both A and R are true and R is correct explanation of A
B) If both A and R are true and R is not the correct explanation of A
C) A is true but R is false
D) Both A and R are false
160. Assertion: Klinefelter's syndrome, Turner's syndrome and Down's syndrome are chromosomal disorders.
Reason: In Klinefelter's syndrome, the females are sterile.
- A) If both Assertion and Reason are true and R is correct explanation of Assertion
B) If both Assertion and Reason are true and R is not the correct explanation of Assertion
C) Assertion is true but Reason is false
D) Both Assertion and Reason are false





161. The protein encoded by the genes cry IAc and cry II Ab control cotton _____.
A) Bollworms
B) Ringworm
C) Flatworm
D) Corn borers
162. RNA interference involves silencing of a specific _____ due to a complementary _____ molecule that binds to and prevents translation of the _____.
A) mRNA, dsDNA, mRNA
B) mRNA, dsRNA, mRNA
C) mRNA, dsRNA, tRNA
D) tRNA, dsRNA, tRNA
163. If the gene isolate from marrow cells producing ADA is introduced into cells at which stage it could be a permanent cure.
A) Early blastula stage
B) Early embryonic stage
C) Late embryonic stage
D) Late morula stage
164. The RNA can be removed by treatment with _____ whereas protein can be removed by the treatment of _____.
A) Ribonuclease, lipase
B) Ribonuclease, protease
C) Deoxyribonuclease lipase
D) Deoxyribonuclease protease
165. The following insect represents a particular phenomenon in England, Identify the correct option.



- A) Genetic drift
B) Natural selection
C) Adaptive radiation
D) Convergent evolution
166. Which ones of the following statements are correct?
(i) Proteins are heteropolymers containing strings of amino acids.
(ii) Biologists describe the protein structure at four levels.
(iii) The first amino acid is also called N – terminal amino acid.
(iv) Only some portions of the protein thread are arranged in the form of a helix.
(v) The long protein chain is also folded upon itself like a hollow woollen ball, giving rise to the tertiary structure.
A) only (ii) and (iii)
B) only (iv) and (v)
C) only (i) and (iv)
D) all of them are correct
167. Which of the following hormones is not a steroidal derivative?
A) Estradiol
B) Epinephrine
C) Cortisol
D) Progesterone





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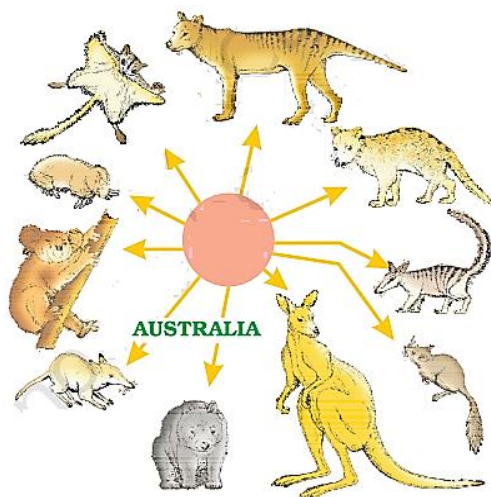
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168. What does the following diagram represent?



A) Natural selection

B) Adaptive radiation of marsupials

C) Placental radiation

D) Mutation

169. Match the columns I and II.

Column - I	Column - II
A. CMI	P. Antibodies are produced outside the body
B. PMNL	Q. Antibodies are produced in the host body
C. Active immunity	R. Cell mediated immunity
D. Passive immunity	S. Leucocytes

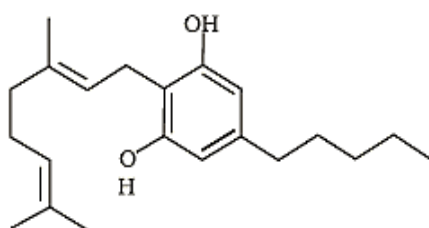
A) A - R, B - S, C - Q, D - P

B) A - S, B - R, C - P, D - Q

C) A - P, B - R, C - Q, D - S

D) A - Q, B - P, C - R, D - S

170. Identify the molecule shown below.



A) Skeletal structure of cannabinoid

B) Chemical structure of marijuana

C) Chemical structure of Morphine

D) Chemical structure of LSD

171. The body in sponges is supported by a skeleton made up of

A) Spicule

B) Spongin fibres

C) Spongocoel

D) Both A & B

172. In frog hind limbs have _____ digits and they are _____ than forelimbs that end with _____ digits

A) 5, larger and muscular, 4

B) 4, larger and muscular, 5

C) 5, smaller and non - muscular, 4

D) 6, larger and neural, 4



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173. Exoskeleton of arthropods, have a complex polysaccharide called _____ and is _____
A) Cellulose, heteropolymer B) Chitin, homopolymer
C) Pectin, homopolymer D) Starch, homopolymer
174. CO₂ is carried by haemoglobin as
A) Carbonic acid B) Oxyhaemoglobin
C) Carbonic anhydrase D) Carbamino – haemoglobin
175. The body of a (d) _____ is covered by a (c) and is unsegmented with a distinct head, muscular foot and (e) _____, over which a (b) _____ and spongy layer of skin forms a (a) _____.
A) (a) Molluscan; (b) Calcareous shell; (c) Visceral hump; (d) Soft; (e) Mantle
B) (c) Molluscan; (b) Calcareous shell; (a) Visceral hump; (d) Soft; (e) Mantle
C) (a) Molluscan; (e) Calcareous shell; (c) Visceral hump; (d) Soft; (b) Mantle
D) (d) Molluscan; (c) Calcareous shell; (e) Visceral hump; (b) Soft; (a) Mantle
176. Which plays role in the maintenance of pH and ionic balance of blood by selective secretion of H⁺ and K⁺ ion.
A) Vasa recta B) Henle's loop
C) Proximal convoluted tubule D) Collecting duct
177. Pons consists of _____ that interconnect different region of the brain
A) Limbic system B) Fibre tract
C) Cerebral aqueduct D) Spinal cord
178. Cortisol helps in maintaining
A) Kidney function B) Cardiovascular system
C) Glucose homeostasis D) Both A & B
179. Fill in the blanks.
- | Year | Population of India (in million) | Population of world (in million) |
|------|----------------------------------|----------------------------------|
| 1990 | 280 | A |
| 1947 | B | 2600 |
| 2000 | C | D |
| 2016 | E | F |
- A) A – 3500, B – 200, C – 1000, D – 6000, E – 1300, F – 7000
B) A – 2000, B – 350, C – 1000, D – 6000, E – 1300, F – 7000
C) A – 7000, B – 200, C – 1000, D – 6000, E – 1300, F – 3500
D) A – 3500, B – 1300, C – 1000, D – 6000, E – 200, F – 7000
180. Select the correct match.
- | | |
|--------------------------------------|----------------|
| A) Ribozyme | Nucleic acid |
| B) F ₂ X Recessive parent | Dihybrid cross |
| C) T. H. Morgan | Transduction |
| D) G. Mendel | Transformation |





1. Ans: B)

$$\begin{aligned}[Q] &= [M^0 L^0 T^0] \\ \Rightarrow [Q] &= [M^0 L^0 T^{-1}] \\ \frac{[P]}{[Q]} &= [M^0 L T^0] \\ [P] &= [M^0 L T^0] [M^0 L^0 T^{-1}] \\ \Rightarrow [PQ] &= [M^0 L T^{-1}] [M^0 L^0 T^{-1}] \\ [PQ] &= [M^0 L T^{-2}]\end{aligned}$$

2. Ans: A) $\frac{\Delta t}{t} = \frac{\Delta T}{T} = \frac{0.2}{50}$

$$\frac{\Delta t}{t} \times 100 = \frac{0.2}{50} \times 100 = 0.4\%$$

3. Ans: D)

$$\begin{aligned}\text{Given } \sqrt{\frac{2h}{g}} - \sqrt{\frac{2\left(\frac{h}{2}\right)}{g}} &= \sqrt{2} - 1 \\ \sqrt{\frac{h}{g}} [\sqrt{2} - 1] &= \sqrt{2} - 1 \\ \Rightarrow h &= 10 \text{ m} \\ \Rightarrow u &= \sqrt{2gh} = \sqrt{2 \times 10 \times 10} \\ u &= 10\sqrt{2} \text{ m/s}\end{aligned}$$

4. Ans: A) $R_{\max} = \frac{u^2}{g} = 50m$

$$H_{\max} = \frac{u^2}{2g} = \frac{50}{2} = 25m$$

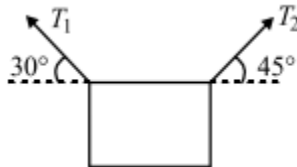
5. Ans: D)

$$\begin{aligned}\vec{A} + \vec{B} &= (2\hat{i} + 3\hat{j}) + (3\hat{i} + 3\hat{j} + 2\hat{k}) \\ &= 5\hat{i} + 6\hat{j} + 2\hat{k} \\ \frac{\vec{A} + \vec{B}}{|\vec{A} + \vec{B}|} &= \frac{5\hat{i} + 6\hat{j} + 2\hat{k}}{\sqrt{5^2 + 6^2 + 2^2}} \\ &= \frac{5\hat{i} + 6\hat{j} + 2\hat{k}}{\sqrt{65}}\end{aligned}$$





6. Ans: B) From F.B.D of block



$$T_1 \cos 30^\circ = T_2 \cos 45^\circ$$

$$\frac{T_1 \sqrt{3}}{2} - \frac{T_2}{\sqrt{2}} = 0$$

$$\sqrt{3}T_1 - \sqrt{2}T_2 = 0$$

7. Ans: D)

$$(f_{\text{lim}})_B = 4 \times 10 \times 0.5 = 20 \text{ N}$$

Since $(f_{\text{lim}})_B > F_{\text{applied}}$

\Rightarrow Block B will not move and tension in the will be zero and block A have no pulling for have no tendency to move.

Hence $f_A = 0$

8. Ans: A)

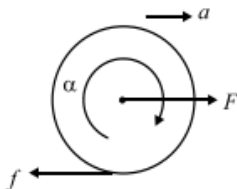
$$W = \vec{F} \cdot \Delta \vec{r}$$

$$W = (\hat{i} + 2\hat{j} + 3\hat{k}) \cdot (3\hat{i} + 4\hat{j} + 5\hat{k})$$

$$= 3 + 8 + 15 = 26 \text{ J}$$

9. Ans: A) Kinetic energy of a moving particle is always positive and it is a scalar quantity.

10. Ans: D)



$$F - f = ma \quad \dots\dots\dots (1)$$

$$f \cdot R = \frac{2}{5} m R^2 \cdot \alpha \quad \dots\dots\dots (2)$$

$$a = R\alpha \quad \dots\dots\dots (3)$$

$$f = \frac{2}{5} ma$$





From equation (1), (2) and (3)

$$F - f = ma$$

$$\Rightarrow F - \frac{2}{5}ma = ma$$

$$\Rightarrow F = \frac{7}{5}ma$$

$$\Rightarrow ma = \frac{5}{7}F$$

$$f = \frac{2}{5}ma = \frac{5}{7}F \times \frac{2}{5} = \frac{2}{7}F$$

11. Ans: D) $|\vec{F}_D|$ and for points A, B and C force of interaction is repulsive and for point E.

Force is attractive in nature. Hence $\vec{F}_C \cdot \vec{F}_E < 0$ because both are in opposite direction.

12. Ans: B) Within elastic limit, elongation is proportional to load because stress \propto strain.

13. Ans: D)

$$\text{For soap bubble, } \Delta P = \frac{4T}{R}$$

$$\frac{\Delta P_1}{\Delta P_2} = \frac{R_2}{R_1} \Rightarrow \frac{R_2}{R_1} = \frac{0.01}{0.02}$$

$$\frac{R_1}{R_2} = 2$$

$$\frac{S_1}{S_2} = \left(\frac{R_1}{R_2}\right)^2 = 4:1$$

14. Ans: B) when temperature of liquid increases then coefficient of viscosity decreases.

15. Ans: C) For anisotropic material

$$\gamma = \alpha_x + \alpha_y + \alpha_z = \alpha + 2\alpha + 2\alpha = 5\alpha$$

16. Ans: A) Area under the curve for black body spectrum is directly proportional to T^4 .

$$\text{Hence } \frac{A_1}{A_2} = \left(\frac{T_1}{T_2}\right)^4 = \left(\frac{T}{2T}\right)^4$$

$$\frac{A_1}{A_2} = \frac{1}{16}$$

17. Ans: B)

Heat is supplied in process DA and AB

$$Q_{in} = n \left(\frac{5R}{2} \right) 1000 + n \left(\frac{7R}{2} \right) 2000$$

$$= \left(\frac{19000}{2} \right) nR$$

$$Q_{rejected} = -n \left(\frac{5R}{2} \right) 2000 - n \left(\frac{7R}{2} \right) \times 1000$$

$$= - \left(\frac{17000}{2} \right) nR$$

$$W = Q_{in} + Q_{rejected}$$





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$$= \frac{19000}{2} nR - \frac{17000}{2} nR = \left(\frac{2000}{2} \right) nR$$

$$\text{Now, efficiency } \eta = \frac{W}{Q_{in}} \Rightarrow \eta = \frac{(2000) \frac{nR}{2}}{19000 \left(\frac{nR}{2} \right)}$$

$$\eta = \frac{2}{19}$$

18. Ans: D)

Given,

$v_{rms} = v$ for diatomic gas

Temperature = T

$v_{rms}^1 = 2v$ for monoatomic gas

Let M = molar mass of diatomic gas. Then, as it dissociates into atoms (monoatomic gas), new molar mass $M' = \frac{M}{2}$. Let T' be the required

temperature.

Now,

rms speed v_{rms} is given as:

$$v_{rms} = \sqrt{\frac{3RT}{M}}$$

$$\text{So, } v = \sqrt{\frac{3RT}{M}}; 2v = \sqrt{\frac{3RT'}{\frac{M}{2}}}$$

$$\text{or, } \boxed{v^2 = \frac{3RT}{M}}; \boxed{4v^2 = \frac{3RT' \times 2}{M}}$$

$$\text{So, } \frac{4v^2}{v^2} = \frac{6RT'}{M} \times \frac{M}{3RT}$$

$$4 = \frac{2T'}{T}$$

$$\Rightarrow \boxed{T' = 2T}$$

19. Ans: C)

When a string is compressed then it does not provide any restoring force.

$$\text{Hence } T = \frac{T_1}{2} + \frac{T_2}{2}$$

$$T = \frac{2\pi}{2} \sqrt{\frac{m}{2k}} + \frac{2\pi}{2} \sqrt{\frac{m}{k}} = \pi \sqrt{\frac{m}{k}} \left[\frac{1}{\sqrt{2}} + 1 \right]$$

$$= (\sqrt{2} + 1) \pi \sqrt{\frac{m}{2k}}$$





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20. Ans: B)

$$f = \frac{v}{2L} = \frac{1}{2L} \sqrt{\frac{T}{\mu}}$$

$$f \propto \sqrt{T}$$

$$f^2 \propto T$$

21. Ans: A)

$$\frac{5V}{2L_B} = \frac{3V}{4L_C} = f$$

$$L_B = \frac{10L_C}{3}$$

$$= \frac{10 \times 18}{3} = 60 \text{ cm}$$

22. Ans: B)

$$E_{\text{origin}} = k \left[\frac{Q_1}{r_1^2} + \frac{Q_2}{r_2^2} + \dots \right]$$

$$= 9 \times 10^9 \times 10^{-6} \left[1 + \frac{1}{2^2} + \frac{1}{4^2} + \dots \right]$$

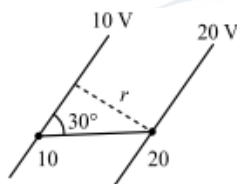
$$= 9 \times 10^3 \times \frac{1}{1 - \frac{1}{4}}$$

$$= 9 \times \frac{4}{3} \times 10^4 = 12000 \text{ N/C}$$

$$F = qE = 2 \times 12000$$

$$= 24000 \text{ N}$$

23. Ans: B)

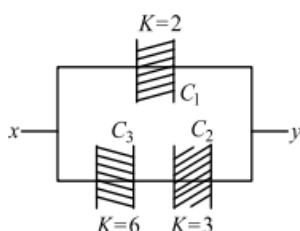


Let any two equipotential surfaces.

$$\text{Then } E = \frac{\Delta V}{r} = \frac{20 - 10}{(0.10 \sin 30^\circ)} = 200 \text{ V/m}$$

Since electric field is normal to equipotential surface and electric field pointed from high potential to low potential.

24. Ans: D) The circuit can be redrawn as





$$C_{xy} = C_1 + \frac{C_1 C_3}{C_2 + C_3} = \frac{2\varepsilon_0 A}{d} + \frac{\frac{3\varepsilon_0 A}{d} \times \frac{6\varepsilon_0 A}{d}}{\frac{9\varepsilon_0 A}{d}} = \frac{2\varepsilon_0 A}{d} + \frac{2\varepsilon_0 A}{d}$$

$$C_{xy} = \frac{4\varepsilon_0 A}{d}$$

25. Ans: C) $R_{eq} = 2 + \frac{2 \times 2}{2+2} + 1 = 4\Omega$

Since no current will flow through 4Ω resistor. Hence $V_{PQ} = \frac{1}{4} \times 12 = 3V$

26. Ans: D) In region BC, current I is decreasing with increase in voltage. Hence this showing the negative resistance region.

27. Ans: C)

$$B = \frac{\mu_0 i}{4\pi d} [\sin \theta_1 + \sin \theta_2]$$

$$= \frac{\mu_0 i}{4\pi d} [\sin \theta_1 + \sin \theta_2]$$

$$= \frac{\mu_0 i}{8\pi d} (\sqrt{3} - 1)$$

28. Ans: B) Magnetic length of a bar magnet is about 0.84 times the geometrical length

29. Ans: B) $F = i l_{eff} B = 2 \times 5 \times 2 = 20 \text{ N}$

30. Ans: C) Eddy current can be minimized by using laminated thin strips of core. Inductor opposes the current change in the electrical circuit. Hence it is called as inertia of the electricity

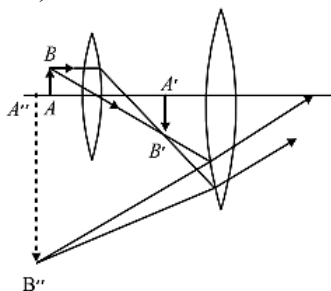
31. Ans: B)

$$\text{Average current} = \frac{\text{Area under curve}}{\text{Time interval}}$$

$$I_{avg} = \frac{\frac{1}{2} \times I_0 \times \frac{T}{2}}{\frac{T}{2}} = \frac{I_0}{2}$$

32. Ans: D) Heat radiations are infrared rays i.e. EM wave. it moves with speed of light in air.

33. Ans: C)





As shown in the ray diagram, intermediate image is real inverted and magnified.

34. Ans: C) Focal length of convex mirror, $f = \frac{R}{2}$, it does not depend on the medium in

which it is placed.

35. Ans: C) $\frac{8\lambda D}{d} = \frac{2\lambda D}{a}$

$$a = \frac{d}{4} = \frac{3}{4} \text{ mm}$$

$$a = 0.75 \text{ mm}$$

36. Ans: C)

$$eV = \frac{hc}{\lambda} - \phi$$

$$\frac{hc}{\lambda} = 3 \text{ eV} + 3.2 \text{ eV}$$

$$= 6.2 \text{ eV}$$

$$\lambda = \frac{12400}{6.2} \text{ \AA}$$

$$\lambda = 2000 \text{ \AA}$$

Hence $\lambda = 2000 \text{ \AA}$ lies in UV region.

37. Ans: A) $E = mc^2 = \frac{10^{-6} \times (3 \times 10^8)^2}{1.6 \times 10^{-19}} \text{ eV} = 5.66 \times 10^{29} \text{ eV}$

38. Ans: B)

For Lyman series

$$\frac{1}{\lambda_L} = R \left(\frac{1}{1} - \frac{1}{n^2} \right)$$

For longest wavelength $n=2$

$$\frac{1}{\lambda_{L, \max}} = R \left(\frac{1}{1} - \frac{1}{2^2} \right)$$

$$\frac{1}{\lambda_{L, \max}} = \frac{3R}{4}$$

For Balmer series

$$\frac{1}{\lambda_B} = R \left(\frac{1}{2^2} - \frac{1}{n^2} \right)$$





For longest wavelength $n = 3$

$$\frac{1}{\lambda_{B, \max}} = R \left(\frac{1}{2^2} - \frac{1}{3^2} \right)$$

$$\frac{1}{\lambda_{B, \max}} = \frac{5R}{36}$$

$$\frac{\lambda_{L, \max}}{\lambda_{B, \max}} = \frac{5R}{36} \times \frac{4}{3R} = \frac{5}{27}$$

$$\frac{\lambda_{L, \max}}{\lambda_{B, \max}} = \frac{5}{27}$$

39. Ans: D) By doping with trivalent atom, semiconductor becomes of p-types i.e. $n_h > n_e$ All types of

semiconductor is electrically neutral.

40. Ans: B) $X = \overline{(A+B)} + \overline{(A+B)} = \overline{(A+B)}$

41. Ans: D)

42. Ans: A)

43. Ans: B)

44. Ans: A) Choke coil offers reactance but a ideal choke coil has a power factor of zero.

45. Ans: B)

CHEMISTRY

46. Ans: A

47. Ans: D) 6.02×10^{23} molecules of CO = 1 mole of CO

6.02×10^{24} molecules of CO = 10 mole of CO

1 mole of CO contains 1 g atom of oxygen

10 mole of CO contains 10 g atom of oxygen

48. Ans: C) Energy values are additives $E = E_1 + E_2$

$$\frac{hc}{\lambda} = \frac{hc}{\lambda_1} + \frac{hc}{\lambda_2}$$

$$\frac{1}{\lambda} = \frac{1}{\lambda_1} + \frac{1}{\lambda_2}$$

By substituting $\lambda = 300 \text{ nm}$, $\lambda_1 = 760 \text{ nm}$

$\therefore \lambda_2 = 495.6 \text{ nm}, \approx 496 \text{ nm}$

49. Ans: D) (A) Electrons in 2s and 2p-orbitals have different screening effect. Hence their Z_{eff} is different.

Z_{eff} off 2s-orbital $>$ Z_{eff} off 2p-orbital.

Therefore, it is not correct.

(B) Energy of 2s-orbital $<$ energy of 2p-orbital.



Hence, it is not correct.

(C) Z_{eff} of 1s-orbital $\neq Z_{\text{eff}}$ of 2s-orbital;

Hence, it is incorrect.

(D) For the two electrons of 2s-orbital, the value of m_s is $+\frac{1}{2}$ and $-\frac{1}{2}$.

Hence, it is correct.

50. Ans: (C) The correct order is $\text{Li} < \text{B} < \text{Be} < \text{C}$. Due to 2s-electrons in B, outer most electron in 2p-orbitals is shielded which results in lowering of ionisation enthalpy

51. Ans: (A) Oxidising property depends on the easiness with which the element is reduced, i.e. it gains electrons which in turn is its electronegativities. Electronegativity $\text{F} > \text{Cl} > \text{O} > \text{N}$ Thus, (A)

52. Ans: D

53. Ans: D

	Lone pair	Hybridisation
	1	sp^3d
	0	sp^3
	2	sp^3d

They have different hybridisation and thus, different structure.

54. Ans: A

	σ	and π	Ratio
A.	8	8	1
B. $\text{O}=\text{C}=\text{O}$	2	2	1
C.	12	3	4
D.	9	2	4.5

Ratio $\text{A} = \text{B} < \text{C} < \text{D}$





55. Ans: C) $W = -p\Delta V = -p(V_2 - V_1) = -1(10 - 5) = -5 \text{ L atm}$

$$= -5 \times 101.27 \text{ J} = -506.35 \text{ J}$$

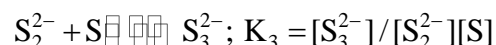
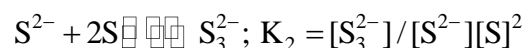
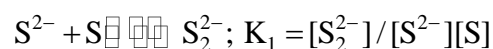
$$q = 400 \text{ J}$$

$$\Delta U = q + W = 400 - 506.35 = -106.35 \text{ J}$$

56. Ans: A) $\Delta G = \Delta H - T\Delta S$

ΔG is negative when $T\Delta S > \Delta H$

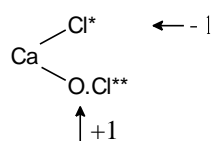
57. Ans: D



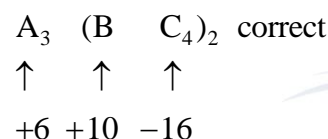
We can prove, $K_1 K_3 = K_2$

$$\therefore K_3 = \frac{K_2}{K_1} = \frac{132}{12} = 11$$

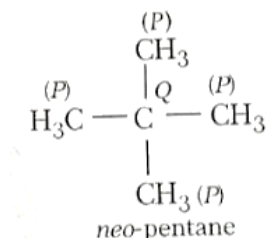
58. Ans: B



59. Ans: B) Compound formed from A, B and C is that for which sum of oxidation number is zero.



60. Ans: C



Here, P = primary carbon atom

Q = quaternary carbon atom

Thus, neo-pentane has four primary carbon atoms, 0-secondary, 0-tertiary and 1-quaternary carbon atoms.

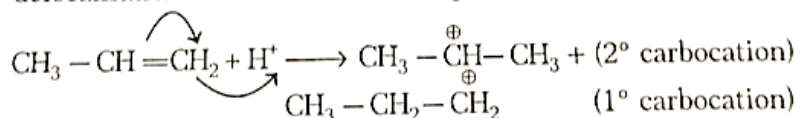
61. Ans: C) There is no additional reagent that can form NO_2^+ for nitration.

Thus, $-\text{CH}_3$ group is oxidised to $-\text{COOH}$.



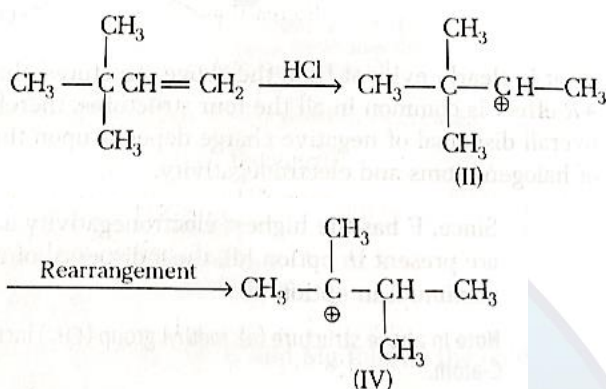


62. Ans: C) When electrophile attacks $\text{CH}_3 - \text{CH} = \text{CH}_2$ then delocalisation of electrons can take place, in two possible ways

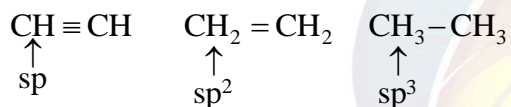


As 2° carbanion is more stable than 1° carbanion, thus first addition is more feasible.

63. Ans: D



64. Ans: A



Electronegativity of carbon $\text{sp} > \text{sp}^2 > \text{sp}^3$

Acidic $\text{CH} \equiv \text{CH} > \text{CH}_2 = \text{CH}_2 > \text{CH}_3 - \text{CH}_3$

Basic $\text{CH} \equiv \overset{\ominus}{\text{C}}(\text{I}) < \text{CH}_2 = \overset{\ominus}{\text{C}}\text{H}(\text{II}) < \text{CH}_3 - \overset{\ominus}{\text{C}}\text{H}_2(\text{III})$

65. Ans: B

66. Ans: B) Methanol has H-bonding. On adding benzene, the molecules of benzene come in between the molecules of methanol. As a result, H-bonds are broken, i.e., intermolecular forces are weakened. Hence, it shows higher vapour pressure than expected from Raoult's law

67. Ans: B

$$\Delta T_b = K_b \times m$$

$$\therefore m = \frac{0.18}{0.512}$$

$$\Delta T_f = K_f \times m = 1.86 \times \frac{0.18}{0.512} = 0.654^\circ$$

$$\therefore T_f = 0^\circ - 0.654^\circ = -0.654^\circ\text{C}$$

68. Ans: C) Greater conductance by HCl is due to greater mobility of H^+ ions.

69. Ans: B





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Cell reaction is



$$E_{\text{cell}} = E^{\circ}_{\text{cell}} - \frac{2 \cdot 303 RT}{n F} \log \frac{[\text{Ti}^{+}]^2}{[\text{Cu}^{2+}]}$$

70. Ans: A

71. Ans: C

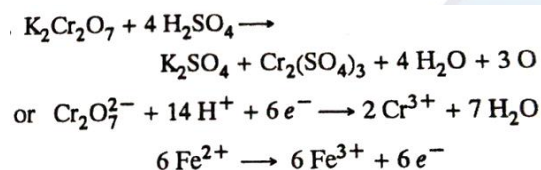
72. Ans: A) Cl^{-} , S^{2-} , P^{3-} are isoelectronic species. Their sizes increase as the nuclear charge decreases from Cl to P.

73. Ans: D) Reducing character of hydrides of group 16 elements increases as the thermal stability decreases from H_2S to H_2Te . H_2O does not behave as a reducing agent. H_2Te is the strongest reducing agent.

74. Ans: A) Electronegativity decreases from F to I and O is next to F in order of decreasing electronegativity.

75. Ans: C) Tungsten, W has the highest melting point (3683 K) among the d-block elements.

76. Ans: A



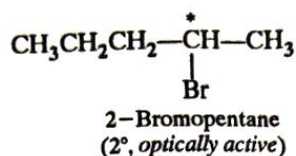
77. Ans: D) In case of lanthanoids, the element with $x f$ electrons has a similar colour to that of $(14 - x)f$ electrons.

78. Ans: B) $[\text{Ni}(\text{CO})_4]$ is tetrahedral with sp^3 hybridization.

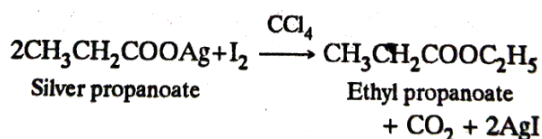
79. Ans: C) $[\text{Fe}(\text{CN})_6]^{3-}$ is paramagnetic but $[\text{Fe}(\text{CN})_6]^{4-}$ is diamagnetic though both are octahedral in shape.

80. Ans: A) Brown ring complex obtained in NO_3^{-} test is due to $[\text{Fe}(\text{H}_2\text{O})_5\text{NO}]^{2+}$

81. Ans: C

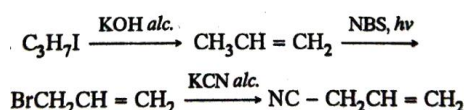


82. Ans: C



This is an example of Birnbaum-Simonini reaction.

83. Ans: C





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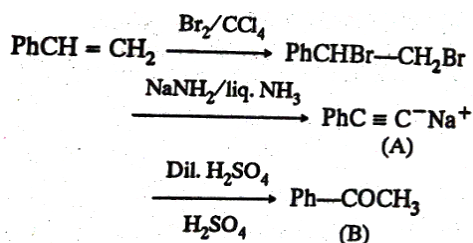
84. Ans: C) (A) CH_3OH and $\text{CH}_2\text{CH}_2\text{CH}_2\text{OH}$ cannot be distinguished because both of them do not react. B) CH_3COCH_3 and $\text{CH}_3\text{CHOHCH}_3$, cannot be distinguished since both of them give iodoform test.

C) $\text{CH}_2\text{CH}_2\text{CH}_2\text{OH}$ and $\text{CH}_3\text{CHOHCH}_3$, can be distinguished since $\text{CH}_3\text{CHOHCH}_3$ gives iodoform test while $\text{CH}_3\text{CH}_2\text{CH}_2\text{OH}$ does not.

D) $\text{CH}_2\text{CH}_2\text{OH}$ and CH_3OH cannot be distinguished since both of them do not respond to iodoform test

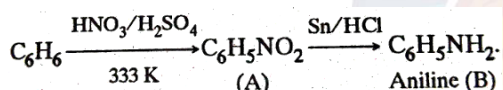
85. Ans: C) Phenol is more acidic than ethanol, methanol and p-methylphenol but is less acidic than o-nitrophenol.

86. Ans: B

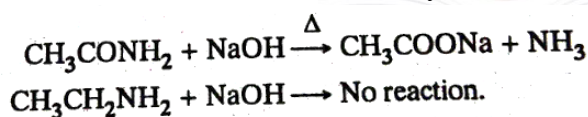


87. Ans: A) As the size of the alkyl groups around CO group increases, crowding increases and hence reactivity decreases, i.e., $\text{CH}_3\text{CHO} > (\text{CH}_3)_2\text{CO} > \text{C}_2\text{H}_5\text{COCH}_3$

88. Ans: B



89. Ans: B) Acetamide and ethylamine can be distinguished by heating with NaOH solution when acetamide evolves NH_3 but ethylamine does not.



90. Ans: B) Guanine is a purine.

BOTANY

91	92	93	94	95	96	97	98	99	100
A	C	D	B	D	D	B	B	A	C
101	102	103	104	105	106	107	108	109	110
B	A	D	C	C	B	A	D	A	A
111	112	113	114	115	116	117	118	119	120
A	B	C	C	C	C	A	A	C	C
121	122	123	124	125	126	127	128	129	130
B	D	A	D	D	A	A	A	B	C
131	132	133	134	135					
D	A	A	B	C					



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136	137	138	139	140	141	142	143	144	145
A	A	B	B	C	C	C	C	D	D
146	147	148	149	150	151	152	153	154	155
B	C	C	A	D	D	A	A	A	A
156	157	158	159	160	161	162	163	164	165
C	A	C	D	C	A	B	B	B	B
166	167	168	169	170	171	172	173	174	175
D	B	B	A	A	D	A	B	D	D
176	177	178	179	180					
D	B	D	B	A					



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