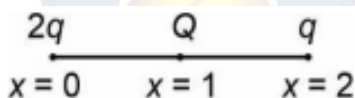
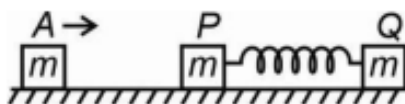




- The ratio of the magnitude of instantaneous velocity and instantaneous speed of a particle is always
A) Less than one B) Greater than one C) Equal to one D) Zero
- In a thermodynamic process, 120 J of heat is added to a gas and work done by the gas is 70 J. The change in internal energy of the gas in the process is
A) 190 J B) - 50 J C) 50 J D) - 70 J
- A particle is moving along x-axis under the influence of a variable force given by $F = \frac{3}{x^2}$, where x is in metre and F is in newton. The work done in displacing the particle from x = 1 m to x = 3 m is
A) 2 J B) 4 J C) $\frac{2}{3}$ J D) $\frac{4}{3}$ J
- Water is flowing through a horizontal tube of non uniform cross-section. At the point of smallest cross-section in the tube, the water will have
A) Maximum speed and maximum pressure
B) Maximum speed and minimum pressure
C) Minimum speed and maximum pressure
D) Minimum speed and minimum pressure
- Three charges are placed in a straight line as shown in the figure. The resultant force on q will be zero if Q is equal to



- A) $\frac{-3q}{2}$ B) $\frac{-q}{2}$ C) $\frac{-2q}{3}$ D) $\frac{q}{2}$
- A parallel plate capacitor has a capacitance 80 μF in air and 140 μF when immersed in oil. The dielectric constant of the oil is
A) 1.5 B) 1.45 C) 2.25 D) 1.75
 - A circuit consisting of a resistance and an inductance coil in series is connected with alternating voltage of 20 V. If the voltage across the resistance is 12 V, then voltage across the coil is
A) 8 V B) 20 V C) 16 V D) 6 V
 - Two identical blocks P and Q each of mass m are connected with a light spring and are kept initially at rest on a smooth surface. A third identical block moving towards positive x-axis strikes block P elastically, then the centre of mass of P and Q

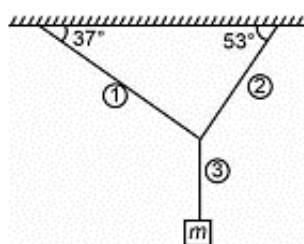


- A) Will move towards +x axis only B) Will move towards -x axis only
C) Will oscillate D) will remain at rest





9. Assertion (A): The focal length of a lens changes when red light is replaced by blue light.
Reason (R): Refractive index of lens depends on the wavelength of light used.
A) Both assertion and reason are true and reason is the correct explanation of assertion.
B) Both assertion and reason are true and reason is not the correct explanation of assertion.
C) Assertion is true but reason is false.
D) Both assertion and reason are false.
10. A block of mass 10 kg is hung with the help of light strings as shown. The tension in wire 1 is [use $g = 10 \text{ m/s}^2$]

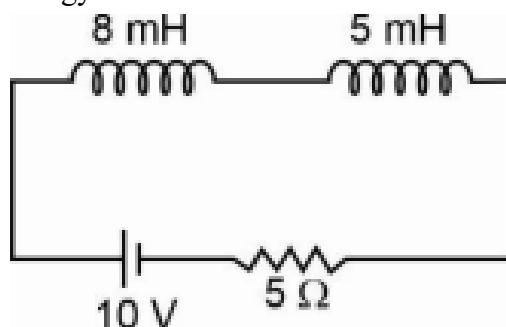


- A) 80 N B) 60 N C) 30 N D) 42 N
11. A small pencil is placed in front of a concave mirror of radius of curvature 20 cm. If the image of the pencil formed by the mirror is inverted, then the distance of the pencil from the mirror may be equal to
A) 20 cm B) 12 cm C) 8 cm D) both A and B
12. The magnetic flux linked with a coil is given by the equation $\phi = 3t^2 + 4t + 1$, where ϕ is in Wb and t is in second. The induced emf in the coil at $t = 2 \text{ s}$ will be
A) 16 V B) 4 V C) 8 V D) Zero
13. When two resistors are connected in parallel, their effective resistance is equal to $\frac{10}{7} \Omega$. If one of the resistors is 2Ω , then resistance of the other resistor is
A) 3Ω B) 5Ω C) $\frac{4}{7} \Omega$ D) $\frac{8}{7} \Omega$
14. For an ideal gas of diatomic molecule (symbols have their usual meanings)
A) $C_v = \frac{3}{2}R$ B) $C_v = \frac{5}{2}R$ C) $C_p = \frac{5}{2}R$ D) $C_p = 3R$
15. Consider the following statements:
(a) Gravitational potential energy of a two point mass system is negative.
(b) In the case of a spherical shell, the plot between gravitation potential and distance from the centre of the shell is continuous.
(c) Gravitational field intensity due to earth increases with increase in altitude.
The correct statements is/are
A) (a) and (b) only B) (b) and (c) only C) (a) and (c) only D) (a), (b) and (c)

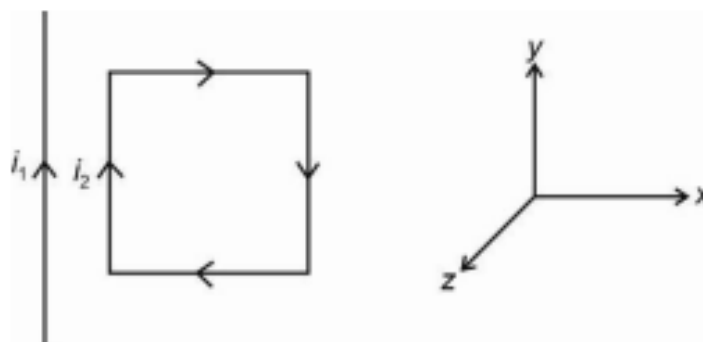




16. Two wires of the same material and same cross section are stretched by the same force. If their lengths are in the ratio 3 : 4 respectively, then their elongations are in the ratio
A) 3 : 4 B) 4 : 3 C) 9 : 16 D) 2 : 3
17. In the given circuit, the energy stored in the 5 mH inductor in steady state will be



- A) 26 mJ B) 50 mJ C) 20 mJ D) 10 mJ
18. A square loop of conducting wire has current flowing through it. A long straight current carrying wire is placed in the place of the loop as shown in the figure. The net force on the square loop will be



- A) Towards $+\hat{i}$ B) Towards $+\hat{k}$ C) Towards $-\hat{i}$ D) Towards $-\hat{j}$
19. A particle is executing simple harmonic motion such that its displacement from mean position is x
[$x < \text{amplitude}$]. The total mechanical energy of the particle is proportional to
A) x B) x^2 C) $\sqrt{x}$ D) x^0
20. If same kinetic energy is provided to electron, proton, neutron and α -particle, the maximum de-Broglie wavelength will be associated with
A) Electron B) Proton C) Neutron D) α -particle
21. The frequency of the first line of Lyman series in hydrogen atom is f . The frequency of the corresponding line emitted by a doubly ionized lithium atom will be
A) $\frac{f}{9}$ B) $9f$ C) $27f$ D) $\frac{f}{3}$
22. The radii of curvature of a biconvex lens are 10 cm and 15 cm. If the refractive index of the lens is $\frac{3}{2}$, then the power of the lens is





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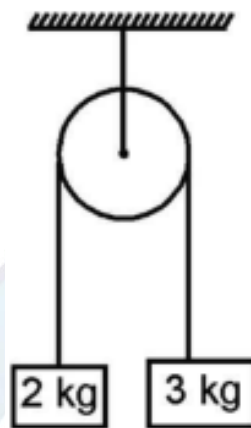
QUESTION PAPER

- A) 1.25 D B) 1.67 D C) 3.2 D D) 8.33 D
23. The binding energies of nuclei A and B are K_1 and K_2 respectively. Four nuclei of A fuse together to give one nuclei of B and energy released is Q. The correct relation of Q with K_1 and K_2 will be
A) $Q = K_1 - K_2$ B) $Q = 4K_1 - K_2$ C) $Q = K_2 - K_1$ D) $Q = K_2 - 4K_1$
24. A light of wavelength $5000 \text{ \AA}$ is incident on a metallic surface of work function 2 eV. The maximum kinetic energy of photoelectrons emitted will be nearly
A) 4.48 eV B) 2 eV C) 0.48 eV D) 2.28 eV
25. Barrier voltage of a p-n junction diode depends on
A) Temperature B) Doping density C) Semiconductor material D) All of these
26. The dimensional formula for a force couple is
A) $[ML^3T^{-2}]$ B) $[MLT^{-3}]$ C) $[ML^2T^{-2}]$ D) $[ML^{-1}T^{-1}]$
27. In an adiabatic process, if pressure of a gas is proportional to the square of its absolute temperature, then the ratio of molar specific heat at constant volume to that at constant pressure is
A) $\frac{1}{2}$ B) $\frac{3}{2}$ C) $\frac{2}{3}$ D) 2
28. A point charge of $1 \mu\text{C}$ is placed at a point A(1, 0, 0)m and another identical charge is placed at B(0, 0, 1) m. The force between them is
A) $18 \times 10^{-3} \text{ N}$ B) $4.5 \times 10^{-3} \text{ N}$ C) $6 \times 10^{-3} \text{ N}$ D) $9 \times 10^{-3} \text{ N}$
29. The output of an AND gate is connected to both the inputs of NOR gate. The combination will act as a
A) NOT gate B) AND gate C) NOR gate D) NAND gate
30. Two particles each of mass 'm' move in a circle of radius '2r' under the action of their mutual gravitational attraction. The speed of each particle will be
A) $\sqrt{\frac{Gm}{r}}$ B) $\frac{1}{2}\sqrt{\frac{Gm}{r}}$ C) $\frac{1}{2}\sqrt{\frac{Gm}{2r}}$ D) $\sqrt{\frac{Gm}{2r}}$
31. Electric field due to a short dipole on its axis depends on distance r from the dipole as
A) r^2 B) $\frac{1}{r^3}$ C) r D) $\frac{1}{r^2}$
32. A parallel plate capacitor with a dielectric between the plates is charged completely and then battery is disconnected, if the dielectric is pulled out, then
A) Charges stored in the capacitor will decrease.
B) Potential difference between the plates will decrease.
C) Capacitance will increase
D) Energy stored in the capacitor will increase.
33. A slit of width d is illuminated by a light of wavelength ($\lambda = 6000 \text{ \AA}$). If the first minima falls at $\theta = 30^\circ$, then value of d will be

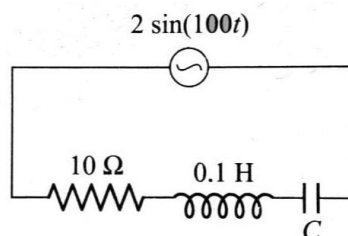




- A) $1.2 \mu\text{m}$ B) $6 \times 10^{-4} \text{ mm}$ C) $3250 \text{ }^\circ\text{A}$ D) $12 \times 10^{-4} \text{ m}$
34. In a certain vernier calipers, 25 divisions on vernier scale have same length as 24 divisions on main scale. One division on main scale is 1 mm long. The least count of the instrument is
- A) 0.04 mm B) 0.01 mm C) 0.02 mm D) 0.08 mm
35. In the given system of masses, the net work done by the tension in the string during the fourth second, after the system is released from rest, will be ($g = 10 \text{ m/s}^2$)



- A) 40 J B) 100 J C) Zero D) 150 J
36. Rate of emission of a black body at 546°C is E. Then the rate of emission of radiation of same body at 0°C will be
- A) $81E$ B) $3E$ C) $\frac{E}{3}$ D) $\frac{E}{81}$
37. Consider the following two statements:
- (a) Electrostatic field is always conservative in nature.
- (b) In electromagnetic waves, the phase difference between oscillating electric and magnetic fields is 90° .
- The correct statement (s) is/are
- A) Only (a) B) Only (b)
- C) Both (a) and (b) D) Neither (a) nor (b)
38. The power factor of the circuit shown in the figure is $\frac{1}{\sqrt{2}}$. The capacitance of the circuit is equal to



- A) $400 \mu\text{F}$ B) $300 \mu\text{F}$ C) $500 \mu\text{F}$ D) $200 \mu\text{F}$





39. Photons of wavelength 660 nm are emitted from a 60 watt lamp. What is the number of photons emitted per second?
A) 2×10^{16} B) 2×10^{28} C) 2×10^{20} D) 2×10^{22}
40. For the astronomical telescope, four lenses of focal length 10 cm, - 10 cm, 100 cm and - 100 cm are available. The focal length of objective lens for maximum magnification should be
A) 10 cm B) - 10 cm C) 100 cm D) - 100 cm
41. The deflection in a galvanometer of resistance 60 Ω , decreases from 20 mA to 4 mA by a shunt of resistance
A) 30 Ω B) 20 Ω C) 15 Ω D) 12 Ω
42. If a parent nucleus ${}_Z X^A$ emits two β positive particles, then for daughter nucleus, atomic number is
A) Z B) $Z - 2$ C) $Z + 1$ D) $Z + 2$
43. In uniform circular motion, which of the following statements is incorrect?
A) Change in magnitude of acceleration is zero.
B) Magnitude of change in velocity is non-zero.
C) Change in magnitude of velocity is zero.
D) Magnitude of change in linear momentum is zero.
44. The equation of a wave pulse travelling along x-axis is given by $y = \frac{30}{2 + (x - 20t)^2}$, x and y are in metre and t is in second. The amplitude of the wave pulse is
A) 5 m B) 20 m C) 15 m D) 30 m
45. Pure silicon at 300 K has equal electron (n_e) and hole (n_h) concentration of $1.5 \times 10^{16} \text{ m}^{-3}$. Doping by indium increases n_h to $4.5 \times 10^{22} \text{ m}^{-3}$. The value of n_e in the doped silicon is
A) $5 \times 10^9 \text{ m}^{-3}$ B) $2.25 \times 10^{11} \text{ m}^{-3}$ C) $9 \times 10^5 \text{ m}^{-3}$ D) $3 \times 10^{19} \text{ m}^{-3}$

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CHEMISTRY

46. Molality of 40% (by mass of solution) urea solution is
A) 0.40 B) 0.60 C) 11.11 D) 0.167
47. Which hydrogen like species will have same radius as that of Bohr orbit of hydrogen atom?
A) $n = 2, \text{Li}^{2+}$ B) $n = 2, \text{Be}^{3+}$ C) $n = 2, \text{He}^+$ D) $n = 3, \text{Li}^{2+}$
48. Wave number of a spectral line for a given transition is $x \text{ cm}^{-1}$ for He^+ , then its value for Be^{3+} (isoelectronic of He^+) for same transition is
A) $x \text{ cm}^{-1}$ B) $4x \text{ cm}^{-1}$ C) $\frac{x}{4} \text{ cm}^{-1}$ D) $2x \text{ cm}^{-1}$
49. The first element of a group in many ways differs from the other heavier members of the group. This is due to
A) the small size





- B) the high electronegativity and high ionization potential
C) the unavailability of d-orbitals
D) All of the above
50. Electronic configuration of four elements A, B, C and D are given below
A. $1s^2 2s^2, 2p^6$ B. $1s^2 2s^2, 2p^4$ C. $1s^2 2s^2, 2p^6 3s^1$ D. $1s^2 2s^2, 2p^5$
Which of the following is the correct order of increasing tendency to gain electron?
A) $A < C < B < D$ B) $A < B < C < D$ C) $D < B < C < A$ D) $D < A < B < C$
51. The formation of oxide ion $O^{2-}(g)$, from oxygen atom requires first an exothermic and then an endothermic step as shown below
 $O(g) + e^- \longrightarrow O^-(g); \Delta H^\ominus = -141 \text{ kJ mol}^{-1}$
 $O^-(g) + e^- \longrightarrow O^{2-}(g); \Delta H^\ominus = +780 \text{ kJ mol}^{-1}$
Thus, process of formation of O^{2-} in gas phase is unfavourable even though O^{2-} is isoelectronic with neon. It is due to the fact that
A) oxygen is more electronegative
B) addition of electron in oxygen results in largest size of the ion
C) electron repulsion outweighs the stability gained by achieving noble gas configuration
D) O^- ion has comparatively smaller size than oxygen atom
52. The increasing order for the solubility of $NaCl$, Na_2SO_4 and Na_3PO_4 in water.
A) $NaCl < Na_2SO_4 < Na_3PO_4$ B) $Na_3PO_4 < Na_2SO_4 < NaCl$
C) $NaCl < Na_3PO_4 < Na_2SO_4$ D) $Na_2SO_4 < NaCl < Na_3PO_4$
53. When N_2^+ is formed from N_2 bond-order and when O_2^+ is formed from O_2 bond-order
A) increases B) decreases C) increases, decreases D) decreases, increases
54. The table shown below gives the dissociation energies (E_{diss}) for single covalent bonds of carbon (C) atoms with element 'A', 'B', 'C' and 'D'. Which element has the smallest atoms?

Bond	$E_{\text{diss}} (\text{kJ mol}^{-1})$
C – A	240
C – B	328
C – C	276
C – D	485

- A) A B) B C) C D) D
55. 1 mole of liquid water vaporises at 1 atmospheric pressure. Given $\Delta H_{\text{vap}}^\ominus = 40.66 \text{ kJ mol}^{-1}$, then internal energy change ($\Delta E^\ominus$) is
A) $37.56 \text{ kJ mol}^{-1}$ B) $-37.56 \text{ kJ mol}^{-1}$ C) $40.66 \text{ kJ mol}^{-1}$ D) $-40.66 \text{ kJ mol}^{-1}$





56. In an adiabatic process, no transfer of heat takes place between system and surroundings. Choose the correct option for free expansion of an ideal gas under adiabatic condition from the following.

A) $q = 0, \Delta T \neq 0, W = 0$

B) $q \neq 0, \Delta T = 0, W = 0$

C) $q = 0, \Delta T = 0, W = 0$

D) $q = 0, \Delta T < 0, W \neq 0$

57. Temperature of 1 mole of a gas is increased by 1° at constant pressure. Work done is

A) R

B) 2R

C) $\frac{R}{2}$

D) 3R

58. $A + 2B \rightleftharpoons 2C + D$, initial concentration of B was 1.5 times that of A, but the equilibrium concentration of A and B are found to be equal. The equilibrium constant for the reaction is

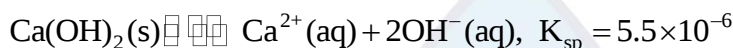
A) 4

B) 8

C) 12

D) 16

59. Slaked lime, Ca(OH)_2 is used extensively in sewage treatment. What is the maximum pH can be established in $\text{Ca(OH)}_2(\text{aq})$?



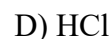
A) 1.66

B) 12.35

C) 7

D) 14

60. Which of the following is the strongest acid?



61. The coefficients w, x, y, z in the reaction $w\text{Cr}_2\text{O}_7^{2-} + x\text{Fe}^{2+} \longrightarrow y\text{Cr}^{3+} + z\text{Fe}^{3+} + \text{H}_2\text{O}$

w x y z

w x y z

w x y z

w x y z

A) 1, 2, 6, 6

B) 6, 1, 2, 4

C) 1, 6, 2, 6

D) 1, 2, 4, 6

62. **Assertion (A):** Among halogens fluorine is the best oxidant.

Reason (R): Fluorine is the most electronegative atom.

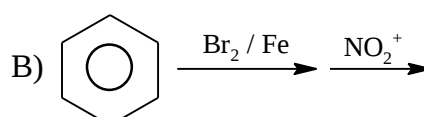
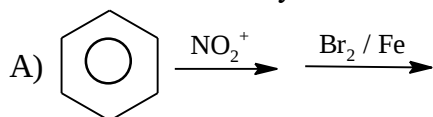
A) Both (A) and (R) are true and (R) is the correct explanation of (A).

B) Both (A) and (R) are true but (R) is not the correct explanation of (A).

C) (A) is true but (R) is false.

D) (A) is false but (R) is true.

63. Which is not the correct synthesis of m-bromo nitrobenzene?



C) Both (A) and (B)

D) None of these

64. $\text{B} \xleftarrow{\text{Lindlar/H}_2} \text{R}-\text{C}\equiv\text{C}-\text{R} \xrightarrow{\text{Na/NH}_3} \text{A}$, A and B are geometrical isomers ($\text{R}-\text{CH}=\text{CH}-\text{R}$)

A) A is cis, B is trans

B) A is trans, B is cis

C) A and B both are cis

D) A and B both are trans





65. Match the reagent from Column I which on reaction with $\text{CH}_3 - \text{CH} = \text{CH}_2$ gives some product given in Column II as per the codes given below

Column I	Column II
A. $\text{O}_3/\text{Zn} + \text{H}_2\text{O}$	1. Acetic acid and CO_2
B. $\text{KMnO}_4 / \text{H}^+$	2. Propan-1-ol
C. $\text{KMnO}_4 / \text{OH}^-$	3. Propan-2-ol
D. $\text{H}_2\text{O} / \text{H}^+$	4. Acetaldehyde and formaldehyde
E. $\text{B}_2\text{H}_6 / \text{NaOH}$ and H_2O_2	5. Propan-1, 2-diol

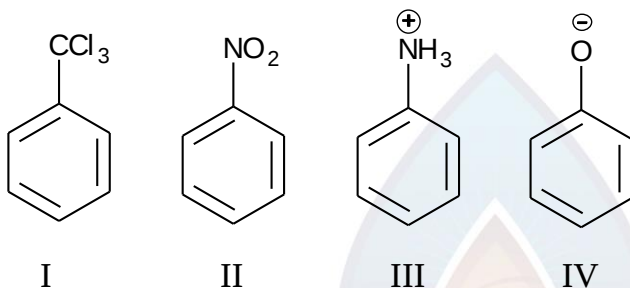
A) A - 4, B - 1, C - 5, D - 3, E - 2

B) A - 1, B - 4, C - 2, D - 3, E - 5

C) A - 4, B - 2, C - 3, D - 5, E - 1

D) A - 5, B - 1, C - 4, D - 3, E - 2

66. Electrophile NO_2^+ attacks the following



In which cases NO_2^+ will be at meta-position?

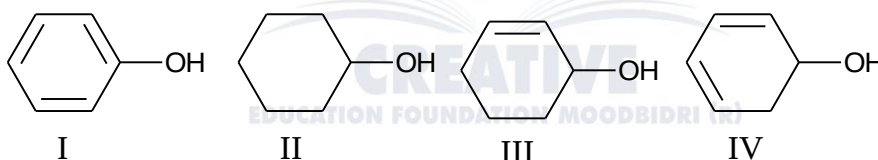
A) Both II and IV

B) I, II and III

C) Both II and III

D) Only I

67. Dehydration of the following in increasing order is



A) $\text{I} < \text{II} < \text{III} < \text{IV}$

B) $\text{II} < \text{III} < \text{IV} < \text{I}$

C) $\text{I} < \text{III} < \text{IV} < \text{II}$

D) $\text{I} < \text{IV} < \text{II} < \text{III}$

68. The elevation in boiling point of a solution of 13.44 g CuCl_2 in 1 kg of water is (Given $K_b = 0.52 \text{ K kg mol}^{-1}$, molecular weight of $\text{CuCl}_2 = 134.4$).

A) 0.05

B) 0.1

C) 0.16

D) 0.21

69. At same temperature, which pair of the following solutions are isotonic solutions?

A) 0.2 M BaCl_2 and 0.2 M urea

B) 0.1 M urea and 0.1 M NaCl

C) 0.1 M NaCl and 0.1 M K_2SO_4

D) 0.1 M $\text{Ba}(\text{NO}_3)_2$ and 0.1 M Na_2SO_4

70. According to Kohlrausch law, the limiting value of molar conductivity of an electrolyte, A_2B is

A) $\lambda_{(\text{A}^+)}^\infty + \lambda_{(\text{B}^-)}^\infty$

B) $\lambda_{(\text{A}^+)}^\infty - \lambda_{(\text{B}^-)}^\infty$

C) $2\lambda_{(\text{A}^+)}^\infty + \frac{1}{2}\lambda_{(\text{B}^-)}^\infty$

D) $2\lambda_{(\text{A}^+)}^\infty + \lambda_{(\text{B}^-)}^\infty$





71. How much quantity has to be passed through 200 ml of 0.5 M CuSO_4 solution to completely deposit copper?
A) 96500 C B) 2×9650 C C) 2×96500 C D) 4×96500 C
72. For the reaction $2A + B \longrightarrow D + E$, the following mechanism has been proposed:
 $A + B \longrightarrow C + D$ (slow)
 $A + C \longrightarrow E$ (fast)
The rate law expression for the reaction is
A) Rate = $k[A]^2[B]$ B) Rate = $k[A][B]$
C) Rate = $k[A][C]$ D) Rate = $k[A]^2[B][C]$
73. An endothermic reaction $A \longrightarrow B$ has an activation energy as $x \text{ kJ mol}^{-1}$ of A. If energy change of the reaction is $y \text{ kJ}$, the activation energy of the reverse reaction is:
A) $-x$ B) $x - y$ C) $x + y$ D) $y - x$
74. Nitrogen forms a variety of compounds in all oxidation states ranging from
A) -3 to $+5$ B) -3 to $+3$ C) -3 to $+4$ D) -3 to $+6$
75. The correct order in which the O – O bond length increases in the following is:
A) $\text{O}_3 < \text{H}_2\text{O}_2 < \text{O}_2$ B) $\text{O}_2 < \text{O}_3 < \text{H}_2\text{O}_2$
C) $\text{O}_2 < \text{H}_2\text{O}_2 < \text{O}_3$ D) $\text{H}_2\text{O}_2 < \text{O}_2 < \text{O}_3$
76. The correct order of increasing bond length of F_2 , N_2 , Cl_2 and O_2 is
A) $\text{O}_2 < \text{N}_2 < \text{Cl}_2 < \text{F}_2$ B) $\text{N}_2 < \text{O}_2 < \text{F}_2 < \text{Cl}_2$
C) $\text{Cl}_2 < \text{O}_2 < \text{N}_2 < \text{F}_2$ D) $\text{F}_2 < \text{Cl}_2 < \text{O}_2 < \text{N}_2$
77. Which of the following is not correct about transition metals?
A) Their melting and boiling points are high
B) Their compounds are generally coloured
C) They can form ionic or covalent compounds
D) They do not exhibit variable valency
78. Number of moles of $\text{K}_2\text{Cr}_2\text{O}_7$ reduced by one mole of Sn^{2+} ion is
A) $\frac{1}{3}$ B) 3 C) $\frac{1}{6}$ D) 6
79. Which of the following ligands is not a chelating agent?
A) EDTA B) en C) Oxalate D) Pyridine
80. The complex ions $[\text{Fe}(\text{CN})_6]^{3-}$ and $[\text{Fe}(\text{CN})_6]^{4-}$
A) Are both octahedral and paramagnetic
B) Are both octahedral and diamagnetic
C) Have same structure but opposite magnetic character
D) Have different structures but same magnetic character.
81. When AgNO_3 is added to a solution of $\text{Co}(\text{NH}_3)_5\text{Cl}_3$, the precipitate of AgCl shows two ionizable chloride ions. This means:
A) Two chlorine atoms satisfy primary valency and one secondary valency.





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- B) One chlorine atom satisfies primary as well as secondary valency
C) Three chlorine atoms satisfy primary valency
D) Three chlorine atoms satisfy secondary valency
82. Among the following secondary allylic halide is
A) 3-Bromo-2-methylbut-1-ene
B) 4-Bromo-3-methylbut-2-ene
C) 3-Bromo-2-methylpropene
D) 1-Bromobut-2-ene
83. $(\text{CH}_3)_3\text{C}-\text{Cl} \xrightarrow[\Delta]{\text{KOH alc.}} \text{A} \xrightarrow[473 \text{ K}]{\text{SO}_2\text{Cl}_2} \text{B}$. The compound B in the above sequence is
A) 3-Chloro-2-methylpropene
B) 1-Chloro-2-methylpropene
C) 1, 2-Dichloro-2-methylpropene
D) 1, 2-Dichloro-2-methylpropene
84. Optical isomers which are non-superimposable mirror images of each other are called
A) Enantiomers
B) Diastereomers
C) Tautomers
D) Geometrical isomers
85. The correct order of relative acidic strength of phenol, ethyl alcohol and water is
A) Phenol > Water > Ethyl alcohol
B) Ethyl alcohol > Water > Phenol
C) Ethyl alcohol > Phenol > Water
D) Water > Phenol > Ethyl alcohol
86. When ethyl iodide is heated with dry silver oxide, it forms
A) Ethyl alcohol
B) Diethyl ether
C) Silver ethoxide
D) Ethyl methyl ether.
87. On heating calcium propionate, the product formed is
A) 3-pentanone
B) 2-Pentanone
C) 3-Methyl-2-butanone
D) Propanone
88. The end product 'C' in the following sequence of chemical reactions is
 $\text{CH}_3\text{COOH} \xrightarrow{\text{CaCO}_3} \text{A} \xrightarrow{\text{heat}} \text{B} \xrightarrow{\text{NH}_2\text{OH}} \text{C}$
A) Acetaldehyde oxime
B) Formaldehyde oxime
C) Methyl nitrate
D) Acetoxime
89. $\text{R}-\text{Cl} + \text{AgCN} \longrightarrow \text{A} \xrightarrow{\text{Redn.}} \text{B}$
A and B respectively
A) RCN , RCH_2NH_2
B) RNC , RNHCH_3
C) RCN , RNHCH_3
D) RNC , RCH_2NH_2
90. Identify 'Z' in the sequence :
 $\text{C}_6\text{H}_5\text{NH}_2 \xrightarrow[273 \text{ K}]{\text{NaNO}_2 + \text{HCl}} \text{X} \xrightarrow{\text{CuCN}} \text{Y} \xrightarrow[\text{Boil}]{\text{H}^+ / \text{H}_2\text{O}} \text{Z}$
A) $\text{C}_6\text{H}_5\text{CN}$
B) $\text{C}_6\text{H}_5\text{CONH}_2$
C) CH_3COOH
D) $\text{C}_6\text{H}_5\text{CH}_2\text{NH}_2$





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BOTANY

QUESTION PAPER

91. The process by which anything is grouped into convenient categories based on some easily observable characters.
A) Nomenclature B) Taxonomy C) Classification D) Identification
92. Linnaeus used _____ as the title of his publication.
A) Systematics B) Systema naturae C) Systema D) Naturae
93. Identify the kingdom having following characters.
a) Cell body contains a well-defined nucleus and other membrane bound organelles
b) Some have flagella or cilia
c) Reproduce asexually and sexually by the process involving cell fusion and zygote formation.
A) Protista B) Monera C) Pteridophyta D) Algae
94. Identify false statement for the kingdom fungi
a) They can reproduce by producing both endo and exogenous spores.
b) Asexual reproduction takes place by zoospore or by aplanospores.
c) They grow in soil, on logs and tree stumps and in living plant bodies as parasites.
d) They prefer to grow in humid and warm places.
A) One B) two C) four D) none
95. Identify the diseases that are caused by viruses.
A) Mumps and herpes B) mumps and cholera
C) Small pox and pneumonia D) Influenza and ring worm
96. Oogamous type of is reproduction observed in
A) *Fucus* and *Ulothrix* B) *Volvox* and *Fucus*
C) *Eudorina* and *Volvox* D) *Fucus* and *Spirogyra*
97. Asexual reproduction in liverworts takes place by fragmentation of thalli or by the formation of specialised structure called
A) Antheridia B) Gemmae C) Prothallus D) Strobili
98. Heterosporous species of the following is
A) *Selaginella* B) *Equisetum* C) *Polytrichum* D) *Marchantia*
99. Which is the site of light reaction
A) Grana in Plastid B) Stroma in plastid
C) Matrix in Mitochondria D) Grana in Mitochondria
100. The interphase nucleus has a loose and indistinct network of nucleoprotein fibre called
A) Chromosome B) chromatin C) centromere D) kinetochore
101. Identify the organelle involved in the following function.
a. Mechanical support
b. motility
c. Maintenance of shape of the cell
A) Nucleus B) Cilia
C) Cytoskeleton D) Plasma membrane





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102. Statement I: the mode of arrangement of sepals or petals in a floral bud with respect to the other members of the same whorl is called phyllotaxy.

Statement II: The margin of the thalamus grows upwardly enclosing the ovary completely and getting fused with it, the other parts of the flower arise above the ovary, hence it is called superior.

- A) Both statement I and statement II are incorrect
- B) Statement I is correct but statement II is incorrect
- C) Statement I is incorrect but statement II is correct
- D) Both statement I and statement II are correct

103. Select the incorrect pair.

- A) Axile placentation – Lemon
- B) Parietal placentation – Argemone
- C) Basal placentation – mustard
- D) Free central placentation – Primrose

104. Cladode is modified

- A) Leaf
- B) stem
- C) root
- D) petiole

105. Assertion: The outside of the epidermis is often covered with a waxy thick layer called cuticle.

Reason: The cuticle prevents loss of water.

- A) Both Assertion and Reason are correct but Reason is not the correct explanation for Assertion
- B) Assertion is correct but Reason is not correct
- C) Assertion is not correct but Reason is correct
- D) Both Assertion and Reason are correct and Reason is the correct explanation for Assertion

106. The xylem and phloem are jointly situated along the same radius of vascular bundle, such vascular bundle is commonly seen in

- A) Leaves
- B) stem
- C) root
- D) both stem and leaves

107. Statement A: Using a prism Julius Von Sachs split light into its spectral components and then illuminated a green alga placed in a suspension of aerobic bacteria.

Statement B : The bacteria were used to detect the sites of O_2 evolution.

- A) Both statement A and statement B are false.
- B) Statement A is true but statement B is false
- C) Statement A is false but statement B is true.
- D) Both statement A and statement B are true.

108. Each photosystem has all the pigments except one molecule of chlorophyll-a, forming a light harvesting system called

- A) Reaction centre
- B) Antennae
- C) Electron transport system
- D) Light harvesting complex

109. The fixation of _____ molecules of CO_2 and _____ turns of cycles are required for the formation of _____ molecule of glucose from C_3 pathway.





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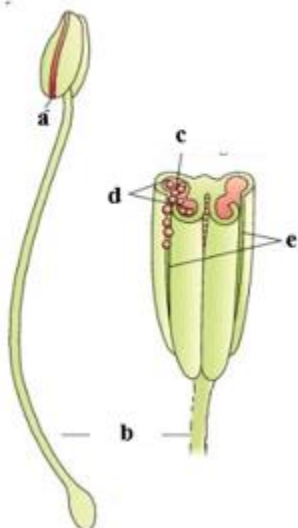
QUESTION PAPER

- A) Six, 5, one
C) six, 6, two
- B) five, 6, two
D) six, 6, one
110. Match the following with their factor and affect for the rate of photosynthesis.
- | Factor | Affect |
|----------------------------------|--|
| 1. Light | P. leaves wilt |
| 2. CO ₂ concentration | Q. breakdown chlorophyll, when it is high |
| 3. Temperature | R. C ₃ plants respond when it is high |
| 4. Water stress | S. C ₄ plant respond when it is high |
- A) 1 – R, 2 – Q, 3 – S, 4 – P
C) 1 – Q, 2 – R, 3 – S, 4 – P
- B) 1 – R, 2 – Q, 3 – P, 4 – S
D) 1 – S, 2 – P, 3 – R, 4 – Q
111. Glucose is phosphorylated to give rise to glucose 6 – phosphate by the activity of the enzyme,
- A) Sucrase
C) Carboxy peptidase
- B) Invertase
D) Hexokinase
112. Plants unlike animals have no special organs for gas exchange, but they have _____ for this purpose.
- A) Stomata and lenticels
B) Mitochondria
C) thylakoid
D) stomata
113. When carbohydrates are used as respiratory substrate, RQ is about
- A) 0.9
B) 0.6
C) 1.0
D) 0.7
114. Identify incorrect match
- | | | |
|---------------------|---|---------------------------|
| A) Cytokinin | - | Adenine derivative |
| B) Absciscic acid | - | Derivative of carotenoids |
| C) Gibberellic acid | - | Kinetin |
| D) Auxin | - | Indole compound |
115. Statement I : Gibberellins plays an important role in seed development, maturation and dormancy.
- Statement II: ABA acts as an antagonist to gibberellins.
- A) Both statement I and statement II are incorrect
B) Statement I is correct but statement II is incorrect
C) Statement I is incorrect but statement II is correct
D) Both statement I and statement II are correct
116. Statement (A): Each cells of sporogenous tissue is a potential to form pollen or microspore mother cell.
- Statement (B): Each cells of the sporogenous tissue is capable of giving rise to a microspore tetrad.
- A) Both statement A and statement B are incorrect
B) Statement A is correct but statement B is incorrect
C) Statement A is incorrect but statement B is correct
D) Both statement A and statement B are correct





117. Recognize the figure and find out the correct matching.



- A) a – anther, b – filament, c – pollen sac, d-pollen grains, e – line of dehiscence
- B) a – filament, b – anther, c- pollen sacs, d – pollen grains, e – line of dehiscence
- C) a – anther, b – filament, c – pollen grains, d- pollen sacs, e – line of dehiscence
- D) a -filament, b – anther, c – pollen grains, d – line of dehiscence, e – pollen sacs

118. Assertion: Pollen grains are well preserved as fossils.

Reason: In the outer layer of pollen grain sporopollenin is absent.

- A) Both Assertion and Reason are correct but Reason is not the correct explanation of Assertion
- B) Assertion is correct but Reason is not correct
- C) Assertion is not correct but Reason is correct
- D) Both Assertion and Reason are correct and Reason is the correct explanation of Assertion

119. Statement A : At the time of Mendel, the nature of the _____ (i) _____ Regulating the pattern of inheritance was not clear.

Statement B : Over next hundred years the nature of the putative genetic material was investigated culminating in the realization that ____ (ii) ____ is the genetic material, at least for the majority of organisms.

- | | | | |
|---------------|----------------|--------------|--------------|
| A) (i)-factor | (ii) – DNA | B) (i)-genes | (ii) – RNA |
| C) (i)-factor | (ii) – protein | D) (i)-genes | (ii) – lipid |

120. Select the incorrect statements about George Gamov from the following.

- A. He was a biochemist
 - B. He argued that since there are only 4 bases and if they have to code for 20 amino acids, The code should constitute a combination of bases.
 - C. He suggested that in order to code for all 20 amino acids, the code should be made up of three nucleotides.
- A) B and C B) A and C C) A, B and C D) only A





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121. Which enzyme was also helpful in polymerising RNA with defined sequence in a template independent manner?

A) Reverse transcriptase

B) Severo ochoa

C) DNA polymerase

D) RNA polymerase

122. Bacterial artificial chromosome (BAC) and Yeast artificial chromosome (YAC) are usually used as

A) Hosts

B) enzymes

C) vectors

D) Proteins

123. Match the following.

Column I : Enzyme		Column II: Product	
P	RNA polymerase – I	i)	tRNA, 5sr RNA, SnRNA
Q	RNA polymerase – II	ii)	Heterogenous nuclear RNA
R	RNA polymerase – III	iii)	rRNAs (28S, 18S, 5.8S)
		iv)	Precursor of hnRNA, small nuclear RNAs

A) P – iii, Q – ii, R – i

B) P – iv, Q – i, R – ii

C) P – iii, Q – iv, R – ii

D) P – iii, Q – ii, R – i

124. Consider the following statement

I. Pollination ensures that the right type of pollen (compatible pollen) always lands on the stigma.

II. Pistil does not have the ability to recognise pollen, whether it is the right type or of the wrong type

III. The ability of the pistil to recognize the pollen followed by its acceptance or rejection is the result of a continuous dialogue between pollen grain and the ovule.

IV. If the pollination is compatible, the pollen grain germinates on the stigma to produce a pollen tube through the intine.

Identify whether the given statements are true or false.

	I	II	III	IV
A)	T	T	F	F
B)	F	F	F	F
C)	T	T	T	T
D)	F	F	T	T

125. A nanometer is

A) One millionth of a meter that is 10^{-9} m

B) One billionth of a meter that is 10^{-9} m

C) One billionth of a meter that is 10^{-6} m

D) One millionth of a meter that is 10^{-6} m

126. Arrange the steps involved in DNA finger printing in a correct order.

A) Isolation of DNA

B) Separation of DNA fragments by electrophoresis

C) Hybridisation using labelled VNTR probe





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- D) Digestion of DNA by restriction endonucleases.
E) detection of hybridized DNA fragments by auto radiography.
F) Transferring (blotting) of separated DNA fragments to synthetic membrane, such as nitrocellulose or nylon.

A) A, D, B, F, C, E

B) A, C, D, B, F, E

C) A, F, C, D, B, E

D) A, B, F, C, D, E

127. Big holes in swiss cheese are made by a

- A) A machine
B) Methanogens
C) The bacterium *Propionibacterium sharmanii* producing a large amount of CO₂
D) Lactobacillus

128. Match the items in column I and column II and choose the correct answer.

Column - I

Column - II

- a) Lady bird
b) Mycorrhiza
c) Dragonfly
d) Biogas

1. *Methanobacterium*
2. Mosquitoes
3. Aphids
4. *Glomus*

Which of the following is the correct option?

	a	b	c	d
A)	1	4	3	2
B)	3	4	2	1
C)	4	1	2	3
D)	2	4	3	1

129. Statement A: Chemical control methods are adopted in agricultural pest control, based on the ability of predators to regulate prey populations.

Statement B: Herbivores and plants appear to be more adversely affected by competition than carnivores.

A) Statement A is correct

B) Statement B is correct

C) Statement A and statement B are correct

D) Statement A and B are wrong

130. The feeding efficiency of one species might be reduced due to the interprising and inhibitory presence of the other species even if resource is abundant is known as

A) Interspecific competition

B) Interference competition

C) Both A and B

D) commensalism

131. Assertion: Biologists are not sure about how many prokaryotic species there might be.

Reason: Many species of microbes are simply not culturable under laboratory conditions.

A) Both Assertion and Reason are correct but Reason is not the correct explanation of Assertion

B) Assertion is correct but Reason is not correct

C) Assertion is not correct but Reason is correct





D) Both Assertion and Reason are correct and Reason is the correct explanation of Assertion

132. The relation between species richness and area on logarithmic scale gives.

A) $\log S = \log S + Z \log A$

B) $\log S = \log Z + C \log A$

C) $\log S = \log C + Z \log A$

D) $\log S = \log A + Z \log C$

133. Primary productivity is expressed in terms of

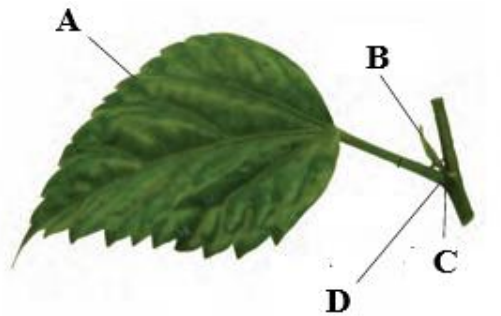
A) Kg/ area

B) weight (g^{-2}) or energy ($Kcal\ m^{-2}$)

C) energy in food calories

D) calories / sq. m

134. Given below is the diagram of a typical leaf. In which of the following all the four parts labelled as A,B, C and D correctly identified.



A

A) Lamina

B) Lamina

C) Lamina

D) Leaflet

B

Axillary bud

Stipule

Axillary bud

Axillary bud

C

Stipule

Axillary bud

Stipule

Stipule

D

leaf base

leaf base

Pedicel

Leaf base

135. Lateral roots arise from

A) Pericycle

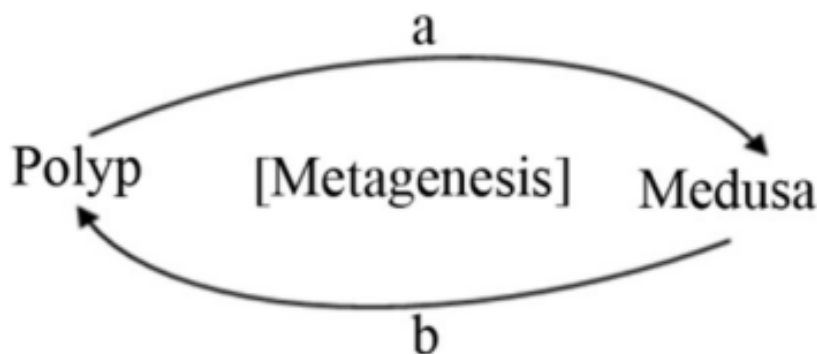
B) Pith

C) stem

D) root

ZOOLOGY

136. Recognise the figure and find out the correct matching.



A) a-sexually, b-asexually

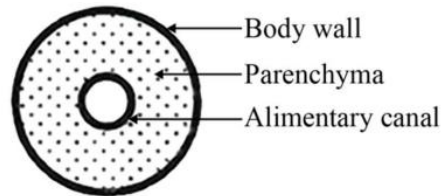
B) a-asexually, b-sexually

C) a-asexually, b-parthenogenetically

D) a-sexually, b-parthenogenetically

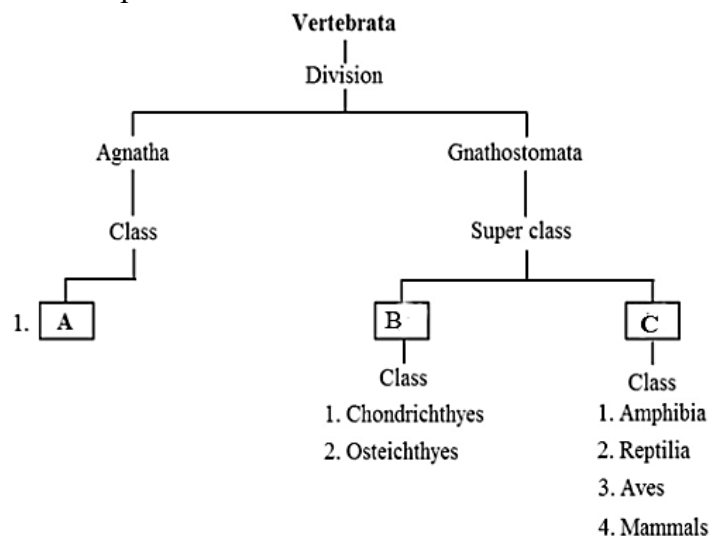


137. The given figure shows a cross section of the body of an invertebrate. Identify the animal which has such a body plan.



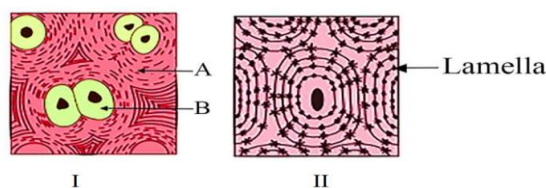
- A) Cockroach (Arthropoda) B) Round worm (Aschelminthes)
C) Planaria (Platyhelminthes) D) Earthworm (Annelida)

138. Go through the following flowchart for division of Subphylum Vertebrata. Fill the gaps A, B, and C select the correct option.



- A) A-Ostracoderm, B-Cyclostomata, C-Pisces
B) A-Cyclostomata, B- Pisces, C-Tetrapoda
C) A-Pisces, B-Tetrapoda, C-Cyclostomata
D) A-Pisces, B-Cyclostomata, C-Tetrapoda

139. Given figures (I and II) show two specialized connective tissues. Identify the figures I, II and the parts labelled as A and B



	I	II	A	B
A)	Cartilage	Bone	Collagen fibers	Chondrocyte
B)	Cartilage	Bone	Collagen fibers	Chondroblast
C)	Bone	Cartilage	Microtubules	Osteoblast
D)	Bone	Cartilage	Collagen fibers	Osteoblast





140. Assertion: Cockroach shows sexual dimorphism.

Reason: The female cockroach bears a pair of short thread like anal styles.

A) both assertion and reason are true, and reason is the correct explanation of assertion.

B) both assertion and reason are true, but reason is not the correct explanation of assertion.

C) assertion is true, but reason is false

D) both assertion and reason are false.

141. Read the statements regarding frogs. Which of the statements is/are correct and incorrect?

(i) The medulla oblongata passes out through the foramen of Magnum and continues into the spinal cord.

(ii) Vasa efferentia are 10-12 in number that arise from testes.

(iii) Ovaries have no functional connection with kidneys.

(iv) Frogs are uricotelic.

A) Statements (i), (ii) and (iii) are correct while statement (iv) is incorrect

B) Statements (i) and (ii) are correct while statements (iii) and (iv) are incorrect

C) Statements (ii) and (iii) are correct while statements (i) and (iv) are incorrect

D) Statements (ii), (iii) and (iv) are correct while statement (i) is incorrect

142. Filtrate obtained after grinding of living tissue is also known as:

A) Slurry

B) Acid - soluble

C) Acid insoluble pool

D) All

143. Which of the following is a Nucleoside?

A) Adenylic acid

B) Uridine

C) Thymidylic acid

D) All

144. Fill up the blanks in the following paragraph by selecting the correct option.

The movement of air into and out of the lungs is carried out by creating a (i) between the lungs and the atmosphere. Inspiration can occur if intra-pulmonary pressure is (ii) than the atmospheric pressure. Expiration takes place when the intra-pulmonary pressure is (iii) than the atmospheric pressure. Inspiration is initiated by the (iv) of diaphragm which (v) the volume of thoracic chamber in the antero-posterior axis.

	(i)	(ii)	(iii)	(iv)	(v)
A)	Concentration gradient	less	higher	relaxation	increase
B)	Concentration gradient	higher	less	contraction	decreases
C)	Pressure gradient	higher	less	relaxation	decreases
D)	Pressure gradient	less	higher	contraction	increases

145. The conditions favourable for dissociation of carbamino haemoglobin are present in

A) alveoli

B) tissues

C) both alveoli and tissues

D) deoxygenated blood

146. Fill in the blanks:

i. Simple organisms like sponges and coelenterates circulate ...1... from their surroundings through their body cavities to facilitate the cells for exchange of O_2 , CO_2 , nutrients and waste products.





ii. More complex organisms use special fluids within their bodies to transport such materials.2... is most commonly used body fluid by most of the higher organisms including humans for this purpose.

iii. Another body fluid...3...also helps in the transport of certain substances.

A) 1-lymph, 2-water, 3-blood

B) 3-lymph, 1-water, 2-blood

C) 2-lymph, 3-water, 1-blood

D) 2-lymph, 1-water, 3-blood

147. When type-B antigen is not present on the surface of RBCs of a person, his plasma would contain _____immunoglobulins.

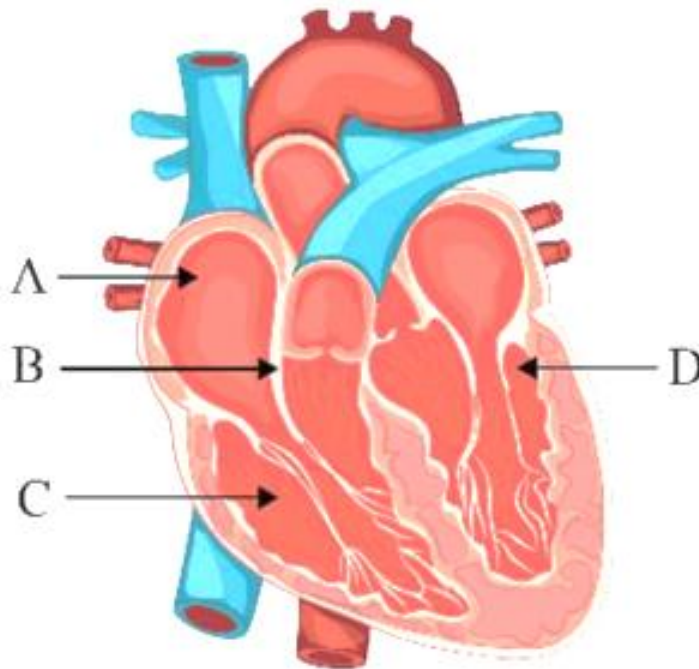
A) Anti-A

B) Anti-B

C) both Antigen-A and Antigen-B

D) none

148. The given figure illustrates a section through the human heart.



Which labelled part represents the site for the generation of action potential in human heart?

A) A

B) B

C) C

D) D

149. Substances like amino acids and glucose in the filtrate are reabsorbed ...a... in the tubular epithelial cells whereas the nitrogenous wastes are absorbed by ...b... transport.

A) a-actively, b-passive

B) a-passively, b-active

C) a-actively, b-active

D) a-passively, b-passive

150. Which factor helps in maintaining an increasing osmolarity towards the inner medullary interstitium?

A. Counter current pattern in vasa recta

B. Counter current pattern in Henle's loop

C. Proximity between the Henle's loop and vasa recta

A) A and B

B) A, B and C

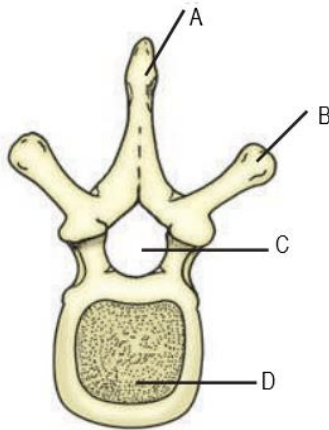
C) C only

D) A and C





151. The following diagram is a typical vertebra of human skeletal system. Select the correct option which chosen the correct labels respectively.



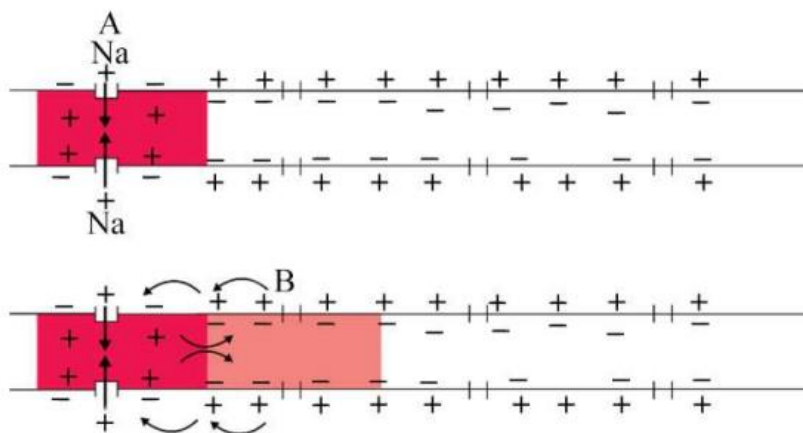
- A) A – centrum; B – neural spine; D – neural canal
- B) A – neural spine; C – neural canal; D - centrum
- C) A – centrum; B – transverse process; C – spinal canal
- D) A – transverse process; C – centrum; D – neural canal

152. Assertion: A person undergo fatigue very soon during exercise.

Reason: Muscle fibres when undergo contraction, the 'I' bands get reduced, whereas the 'A' bands retain its length.

- A) assertion and reason are true, and reason is the correction explanation of assertion.
- B) assertion and reason are true, but reason is not correct explanation of assertion.
- C) assertion is true, but reason is false
- D) both assertion and reason are false

153. During nerve impulse conduction, current flows



- A) From site A to site B on the inner surface
- B) From site B to site A on the outer surface
- C) Both A and B
- D) None





154. Which of the following represents the covering layers of the brain correctly from inside to outside?

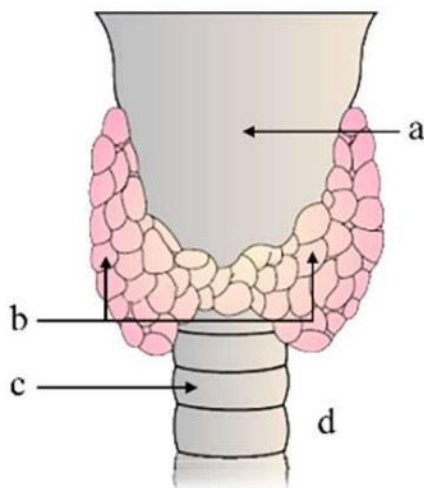
- A) Pia mater - arachnoid mater - dura mater - skull
- B) Dura mater - arachnoid mater - pia mater - meninges - skull
- C) Skull - arachnoid mater - dura mater - pia mater - cerebral meninges
- D) Skull - dura mater - arachnoid mater - pia mater

155. Assertion: Medulla contains centres which control the respiration, cardiovascular reflexes and gastric secretions.

Reason: Medulla contains neurosecretory cells which secretes hormones.

- A) both assertion and reason are true, and reason is correct explanation of assertion.
- B) both assertion and reason are true, but reason is not correct explanation of assertion.
- C) assertion is true, but reason is false
- D) both assertion and reason are false.

156. Recognize the figure and find out the correct matching.



- A) a-isthmus, b-parathyroid, c-thyroid, d-dorsal side
- B) a-vocal cord, b-thyroid, c-trachea, d-dorsal side
- C) a-vocal cord, b-thyroid, c-trachea, d-ventral side
- D) a-vocal cord, b-parathyroid, c-trachea, d-ventral side

157. Match the columns I and II and choose the correct combination from the options given.

Column I

- i. Protein hormone
- ii. Steroid hormone
- iii. Amino acid derivative
- iv Iodothyronines

Column II

- 1. ACTH
- 2. Cortisol
- 3 Thyroxine
- 4. Estradiol
- 5. Epinephrine
- 6. Progesterone
- 7. GnRH
- 8. Somatostatin





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QUESTION PAPER

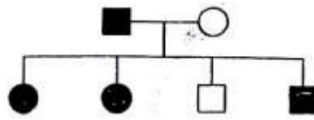
- A) i-7 and 8; ii -4 and 6; iii-1 and 5; iv-3 and 2
B) i-1, 7 and 8; ii-2, 4 and 6; iii-5; iv-3
C) i-2, 4 and 6; ii-2, 7 and 8; iii-3; iv-5
D) i-1, 5 and 6; ii-2, 7 and 8; iii-5; iv-3
158. Which one of the following is the correct matching of the events occurring during menstrual cycle?
- A) Development of corpus luteum→Secretory phase and increased secretion of progesterone.
B) Menstruation→Breakdown of myometrium and ovum not fertilized.
C) Ovulation→LH and FSH attain peak level and sharp fall in the secretion of progesterone.
D) Proliferative phase→Rapid degeneration of myometrium and maturation of Graafian follicle.
159. Which one of the following is the most likely root cause why menstruation is not taking place in regularly cycling human female?
- A) Maintenance of high concentration of sex-hormones in the blood stream
B) Retention of well-developed corpus luteum
C) Fertilisation of the ovum
D) Maintenance of the hypertrophic endometrial lining
160. What happens during fertilisation in humans after many sperms reach close to the ovum?
- A) Secretions of acrosome helps one sperm enter cytoplasm of ovum through zona pellucida
B) All sperms except the one nearest to the ovum lose their tails
C) Cells of corona radiata trap all the sperms except one
D) Only two sperms nearest the ovum penetrate zona pellucida
161. In vitro fertilization is a technique that involves transfer of which one of the following into the fallopian tube?
- A) Zygote only
B) Embryo only, up to 8 cell stage
C) Either zygote or early embryo up to 8 cell stage
D) Embryo of 32 cell stage
162. The technique called gamete intrafallopian transfer (GIFT) is recommended for those females:
- A) Who cannot produce an ovum
B) Who cannot retain the foetus inside uterus
C) Whose cervical canal is too narrow to allow passage for the sperms
D) Who cannot provide suitable environment for fertilisation
163. In case of a couple where the male is having a problem of inability to ejaculate/inseminate, which technique will be suitable for fertilization.
- A) Intrauterine transfer
B) Gamete intracytoplasmic fallopian transfer
C) Artificial Insemination





D) Intracytoplasmic sperm injection

164. Study the pedigree chart of a certain family given below and select the correct conclusion which can be drawn for the character.



- A) The female parent is heterozygous
B) The parents could have had a normal daughter for this character
C) The trait under study could not be colour blindness
D) The male parent is homozygous dominant
165. The F₂ generation in a Mendelian cross showed that both genotypic and phenotypic ratios are same as 1 : 2 : 1. It represents a case of:
A) Dihybrid cross
B) Monohybrid cross with complete dominance
C) Monohybrid cross with incomplete dominance
D) Co-dominance
166. If both parents are carriers for thalassemia, which is an autosomal recessive disorder, what are the chances of pregnancy resulting in an affected child?
A) 25% B) 100% C) no chance D) 50%
167. A human female with Turner's syndrome:
A) Is able to produce children with normal husband
B) Has 45 chromosomes with XO
C) Has one additional X chromosome
D) Exhibits male characters
168. The eye of octopus and eye of cat show different patterns of structure, yet they perform similar function. This is an example of:
A) Analogous organs that have evolved due to convergent evolution.
B) Analogous organs that have evolved due to divergent evolution.
C) Homologous organs that have evolved due to convergent evolution.
D) Homologous organs that have evolved due to divergent evolution.
169. A population will not exist in Hardy-Weinberg equilibrium if:
A) There are no mutations B) There is no migration
C) The population is large D) Individuals selected by nature
170. Artificial selection of low yield cows to obtain cows yielding high milk output represents:
A) Stabilizing selection as it stabilizes a character in the population.
B) Directional as it pushes the mean of a character in one direction
C) Disruptive as it splits the population into the one yielding higher output and the other lower output
D) Stabilizing followed by disruptive as stabilizes the population to produce high yielding cows.





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171. At which stage of HIV infection does one usually show symptoms of AIDS?
- A) When the infecting retrovirus enters host cells
 - B) When viral DNA is produced by using reverse transcriptase
 - C) When HIV replicates rapidly in helper T-lymphocytes and damages large number of these
 - D) Within 15 days of sexual contact with an infected person

172. In which one of the following options the two examples are correctly matched with their immunity?

Examples	Type of Immunity
A) Anti-tetanus and anti-snake bite injections	Active immunity
B) Saliva in mouth and Tears in eyes	Physical barriers
C) Mucus coating of epithelium lining the urinogenital tract and the HCl in stomach	Physiological barriers
D) Polymorphonuclear leukocytes and monocytes	Cellular barriers

173. Infection of Ascaris usually occurs by:

- A) Tse-tse fly.
- B) mosquito bite.
- C) drinking water containing eggs of Ascaris.
- D) eating imperfectly cooked pork.

174. The convention for naming restriction endonucleases is

- A) First two letters come from genus & third from species of prokaryotic cell from which they were isolated.
- B) First two letters come from species & third from genus of prokaryotic cell from which they were isolated.
- C) First letter come from genus & second two from species of prokaryotic cell from which they were isolated.
- D) First letter come from species & second two from genus of prokaryotic cell from which they were isolated

175. State True(T) or False(F).

- A. Ti plasmid is tumour inducing plasmid of Agrobacterium.
- B. Two selectable markers present in pBR322 are Lac 'Z' and amp^R.
- C. Restriction digestion with EcoRI produces blunt ends fragments.
- D. Selection of transformants from non-transformants can be done by plating on agar rich with tetracycline, if insertional inactivation takes in Pvu-I recognition site.

A B C D

- A) F T F F
- C) T F T T

- B) F F T F
- D) T F F T





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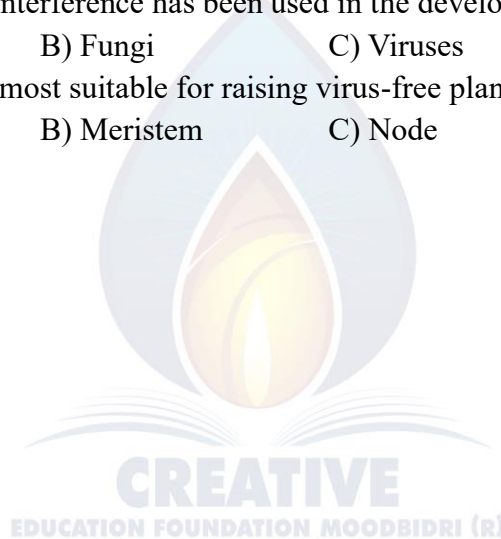
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176. In genetic engineering, the antibiotic-resistant genes are used:
- A) As sequences from where replication starts
 - B) To keep the cultures free of infection
 - C) As selectable markers
 - D) To select healthy vectors
177. Which one of the following is used as vector for cloning genes into animal cells?
- A) Retrovirus
 - B) Baculovirus
 - C) Salmonella typhimurium
 - D) Rhizopus nigricans
178. Genetic engineering has been successfully used for producing:
- A) Animals like bulls for farm work as they have superpower
 - B) Transgenic mice for testing safety of polio vaccine before use in humans
 - C) Transgenic models for studying new treatments for certain cardiac diseases
 - D) Transgenic Cow-Rosie which produces high fat milk for making ghee
179. The process of RNA interference has been used in the development of plants resistant to:
- A) Nematodes
 - B) Fungi
 - C) Viruses
 - D) Insects
180. Which part would be most suitable for raising virus-free plants for micropropagation?
- A) Vascular tissue
 - B) Meristem
 - C) Node
 - D) Bark





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PHYSICS

SOLUTIONS

1. Ans: C)

2. Ans: C) $\Delta Q = \Delta U + \Delta W$

$$120 = \Delta U + 70$$

3. Ans: A) $W = \int_1^3 F dx = \int_1^3 \frac{3}{x^2} dx = 2J$

4. Ans: B) Less A, more speed, more KE, less pressure.

5. Ans: B) $F_1 = F_2$

$$\frac{K2q.q}{2^2} = \frac{KQq}{1^2}$$

$$Q = \frac{-q}{2}$$

6. Ans: D) $K = \frac{C'}{C} = \frac{140}{80} = 1.75$

7. Ans: C) $V = \sqrt{V_R^2 + V_L^2}$

$$V_R = 12V$$

$$V = 20V \Rightarrow V_L = 16V$$

8. Ans: A) To conserve momentum, COM moves +x direction.

9. Ans: A) $\mu = A + \frac{B}{\lambda^2}$

10. Ans: B) $T_3 = 100 \text{ N} = mg$

$$T_2 \sin 53 + T_1 \sin 37 = T_3 \quad \text{--- (1)}$$

$$T_1 \cos 37 = T_2 \cos 53 \quad \text{---- (2)}$$

Solve (1) and (2)

11. Ans: D) $|u| > f$, then inverted image.

12. Ans: A) $e = \frac{d\phi}{dt} = 6t + 4$

$$t = 2s \Rightarrow e = 16v$$

13. Ans: B) $\frac{R_1 R_2}{R_1 + R_2} = \frac{10}{7}$

$$\frac{2R_2}{2 + R_2} = \frac{10}{7}$$

$$R_2 = 5$$





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SOLUTIONS

14. Ans: B)

15. Ans: A)

16. Ans: A)

$$F = \frac{AY}{l_1 \Delta l_1} = \frac{AY}{l_2 \Delta l_1}$$
$$\Rightarrow \frac{\Delta l_1}{\Delta l_2} = \frac{3}{4}$$

17. Ans: D) $L = 5 \times 10^{-3} H$

$$I = \frac{10}{5} = 2A$$

$$E = \frac{1}{2} LI^2 = 10mJ$$

18. Ans: C) The wires have same direction of current are closer. So, net force is attractive.

19. Ans: D) $TE = \text{constant}$.

20. Ans: A) $\lambda = \frac{h}{\sqrt{2mKE}}$

$$\lambda \propto \frac{1}{\sqrt{m}}$$

21. Ans: B) $E = h\delta = 13.6z^2 \left(\frac{1}{n_1^2} - \frac{1}{n_2^2} \right)$

$$f \propto z^2$$

22. Ans: D) $R_1 = +10\text{cm}$, $R_2 = -15\text{cm}$, $\mu_2 = \frac{3}{2}$, $\mu_1 = 1$

$$p = \frac{\mu_2 - \mu_1}{\mu_1} \left(\frac{1}{R_1} - \frac{1}{R_2} \right) = 8.33D$$

23. Ans: D) $Q = (BE)_{\text{products}} - Q(BE)_{\text{reactants}} = K_2 - 4K_1$

24. Ans: C) $E_{\text{photon}} = W_0 + KE_{\text{max}}$

$$KE_{\text{max}} = \frac{12400}{5000} - 2eV = 0.47eV$$

25. Ans: D)

26. Ans: C) $\tau = [ML^2T^{-2}]$

27. Ans: A) $P \propto T^2$





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$$PT^{-2} = \text{constant} \Rightarrow P^{1-\delta} = \text{constant} \quad \text{--- (1)}$$

$$P'T^{\frac{\delta}{1-\delta}} = \text{constant} \quad \text{--- (2)}$$

After comparing (1) and (2)

$$\delta = 2 \quad \frac{C_V}{C_P} = \frac{1}{2}$$

28. Ans: B) $q_1 = q_2 = 1\mu C$

$$r = \sqrt{(1-0)^2 + (0-0)^2 + (0-1)^2}$$

$$F = \frac{kq_1q_2}{r^2} = \frac{9 \times 10^9 \times 10^{-6} \times 10^{-6}}{(\sqrt{2})^2} = 4.5 \times 10^{-3} N$$

29. Ans: D) $(\overline{AB} + AB) = \overline{AB} \Rightarrow \text{NAND GATE}$

30. Ans: B) $\frac{mv^2}{2r} = \frac{Gmm}{(4r)^2} \Rightarrow v = \frac{1}{2} \sqrt{\frac{Gm}{2r}}$

31. Ans: B) $E = \frac{1}{4\pi\epsilon_0} \frac{2p}{r^3}$

32. Ans: D) Q remains same

$$E = \frac{Q^2}{2C}$$

$$C \downarrow, E \uparrow$$

33. Ans: A) $\lambda = d \sin \theta$

$$\lambda = d \sin 30 \Rightarrow d = 1.2\mu m$$

34. Ans: A) 25 VSD = 24 MSD

$$1VSD = \frac{24}{25} MSD$$

$$LC = 1MSD - VSD$$

$$= 1MSD \left(1 - \frac{24}{25} \right) = 1mm \left(1 - \frac{24}{25} \right) = 0.04mm$$

35. Ans: C) Work done by both tension are equal but opposite sign.

36. Ans: D) $E \propto (546 + 273)^4$

$$E' \propto (273)^4$$

$$E' = \frac{E}{81}$$





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SOLUTIONS

37. Ans: A)

$$38. \text{ Ans: C) } \cos \phi = R / Z = \frac{1}{\sqrt{2}} = \frac{10}{\sqrt{10^2 + \left(\frac{50}{\pi} \times 0.1 - \frac{1}{C \times 50 / \pi} \right)^2}}$$

$$39. \text{ Ans: C) } p = \frac{Nhc}{\Delta t \lambda}$$

$$\frac{N}{\Delta t} = \frac{p \lambda}{hc} = 2 \times 10^{20}$$

40. Ans: C)

$$41. \text{ Ans: C) } 4\text{mA} \times 60 = 16 \text{ mA} \times S$$

$$S = 15 \Omega$$

42. Ans: B) In β^+ decay, z decreases by 1.

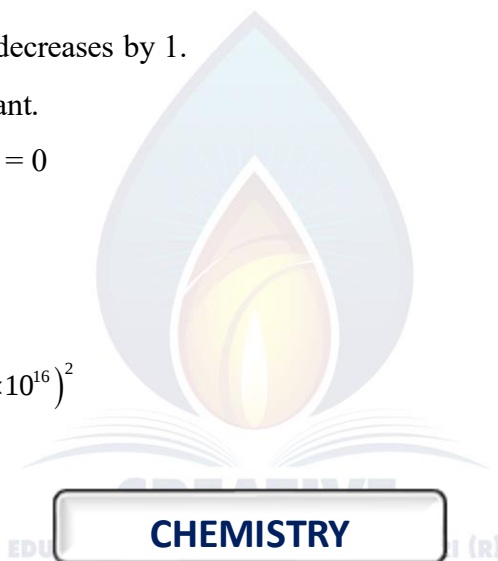
43. Ans: D) Speed is constant.

44. Ans: C) When $x - 20y = 0$

$$y = \frac{30}{2} = 15m = A$$

$$45. \text{ Ans: A) } n_e \times n_h = n_i^2$$

$$n_e \times 4.5 \times 10^{22} = (1.5 \times 10^{16})^2$$



46. Ans: C) 100 g solution has urea = 40 g

$$\therefore 60 \text{ g solution has urea} = 40\text{g} = \frac{40}{60} \text{ mol urea}$$

$$\text{Molality} = \frac{\text{mol}}{\text{kg of solvent}} = \frac{40/60}{0.060} = 11.11$$

$$47. \text{ Ans: B) } r_n = \frac{n^2 a_0}{Z}, n = 2, Z = 4; r_2 = \frac{4a_0}{4} = a_0$$

$$48. \text{ Ans: B) } \bar{\nu}(\text{wave number}) = R_H Z^2 \left[\frac{1}{n_1^2} - \frac{1}{n_2^2} \right]$$





$$\bar{v}_1(\text{He}^+, Z = 2) = R_H(2)^2 \left[\frac{1}{n_1^2} - \frac{1}{n_2^2} \right]$$

$$\bar{v}_2(\text{Be}^{3+}, Z = 4) = R_H(4)^2 \left[\frac{1}{n_1^2} - \frac{1}{n_2^2} \right]$$

$$\frac{\bar{v}_2}{\bar{v}_1} = 4$$

$$\bar{v}_2 = 4\bar{v}_1 = 4x$$

49. Ans: D

50. Ans: (A) Electronic configuration of elements indicates that A is a noble gas (i.e. Ne), B is oxygen (group 16), C is sodium metal (group 1) and D is fluorine (group 17).

(i) Noble gases have no tendency to gain electrons since all their orbitals are completely filled. Thus, element A has the least electron gain enthalpy.

(ii) Since, element D has one electron less and element B has two electrons less than the corresponding noble gas configuration, hence, element D has the highest electron gain enthalpy followed by element B.

(iii) Since, element C has one electron in the s-orbital and hence needs one more electron to complete it, therefore, electron gain enthalpy of C is less than that of element B. Combining all the facts given above, the electron gain enthalpies of the four elements increase in the order $A < C < B < D$.

51. Ans: (C) Although O^{2-} has noble gas configuration isoelectronic with neon but its formation is unfavourable due to the strong electronic repulsion between the negatively charged O^- ion and the second electron being added. Hence, the electron repulsion outweighs the stability gained by achieving noble gas configuration.

52. Ans: B) Larger the ionic character, larger the solubility in water. Smaller the cation, larger the anion, larger the charge then larger the covalent nature and thus smaller the solubility.

When cation is identical, then order is on the basis of anions

Charge of anion (Polarizing power) $\text{Cl}^- < \text{SO}_4^{2-} < \text{PO}_4^{3-}$

Covalent nature $\text{NaCl} < \text{Na}_2\text{SO}_4 < \text{Na}_3\text{PO}_4$





Solubility $\xrightarrow{\text{Na}_3\text{PO}_4 < \text{Na}_2\text{SO}_4 < \text{NaCl}}$

53. Ans: D) Bond order

$\begin{array}{cc} \text{N}_2 & 3.0 \\ \text{N}_2^+ & 2.5 \end{array} \left[\begin{array}{c} \text{decreases} \end{array} \right]$
 $\begin{array}{cc} \text{O}_2 & 2.0 \\ \text{O}_2^+ & 2.5 \end{array} \left[\begin{array}{c} \text{increases} \end{array} \right]$

54. Ans: D) Smallest the atom, larger the overlapping hence higher the bond dissociation energy.

55. Ans: A) $\text{H}_2\text{O}(\text{l}) \rightleftharpoons \text{H}_2\text{O}(\text{s})$

$$\Delta n_g (\text{gaseous species}) = 1 - 0 = 1$$

$$\Delta H^\circ = 40.66 \text{ kJ mol}^{-1}$$

At 1 atmosphere, boiling point = $100^\circ = 373 \text{ K}$

$$\therefore \Delta H^\circ = \Delta E^\circ + \Delta n_g RT$$

$$\therefore \Delta E^\circ = \Delta H^\circ - \Delta n_g RT$$

$$= 40.66 - 1 \times 8.314 \times 10^{-3} \times 373 = 37.56 \text{ kJ mol}^{-1}$$

56. Ans: C) Free expansion, $W = 0$

Adiabatic process, $q = 0$

$\Delta U = q + W = 0$, this means that internal energy remains constant. Therefore, $\Delta T = 0$ in ideal gas there is no intermolecular attraction.

Hence, when such a gas expands under adiabatic conditions into a vacuum, no heat is absorbed or evolved since no external work is done to separate the molecules.

57. Ans: A) $pV = RT$ at temperature TK

$$p(V + \Delta V) = R(T + 1) \text{ at temperature } (T + 1)K$$

$$\therefore pV + p\Delta V = RT + R$$

$$\therefore p\Delta V = R$$

58. Ans: A) $\text{A} + 2\text{B} \rightleftharpoons 2\text{C} + \text{D}$

Initial $1\text{M} \quad 1.5\text{M} \quad 0 \quad 0$

Equili. $(1 - x)\text{M} \quad (1.5 - 2x)\text{M} \quad 2x \text{M} \quad x\text{M}$

$$\text{Given, } 1 - x = 1.5 - 2x \Rightarrow x = 0.5$$

$$\therefore [\text{A}] = 0.5\text{M}, [\text{C}] = 1 \text{M},$$

$$[\text{B}] = 0.5 \text{M}, [\text{D}] = 0.5 \text{M}$$





$$K_c = \frac{[C]^2[D]}{[A][B]^2} = \frac{(1)^2(0.5)}{(0.5)(0.5)^2} = 4$$

59. Ans: B) $K_{sp} = 5.5 \times 10^{-6}$



$$[\text{Ca}^{2+}][\text{OH}^-]^2 = 5.5 \times 10^{-6}$$

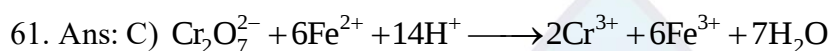
$$(S)(2S)^2 = 5.5 \times 10^{-6}$$

$$S = 0.0111$$

$$\therefore [\text{OH}^-] = 2S = 0.0111 \times 2 = 0.0222$$

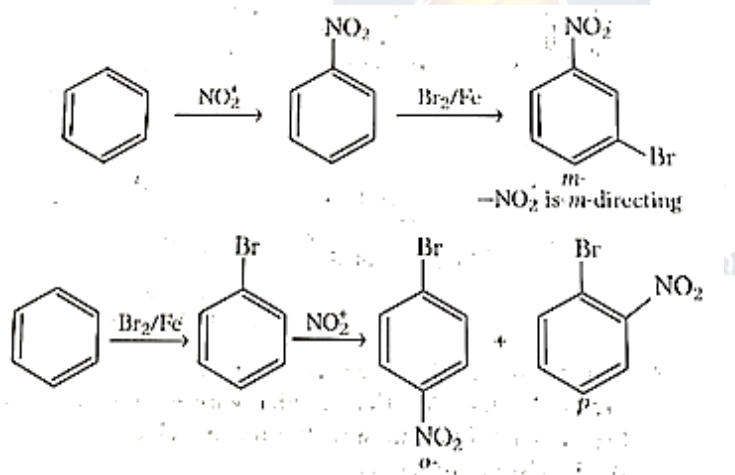
$$\therefore \text{pOH} = 1.65; \text{pH} = 12.35$$

60. Ans: A) Maximum oxidation number of the central atom is in HClO_4 (= 7)



62. Ans: B) Both Assertion and Reason are true but Reason is not the correct explanation of Assertion. Among halogen F_2 is the best oxidant because it has the highest E° value.

63. Ans: B



64. Ans: B

65. Ans: A) A - 4, B - 1, C - 5, D - 3, E - 2





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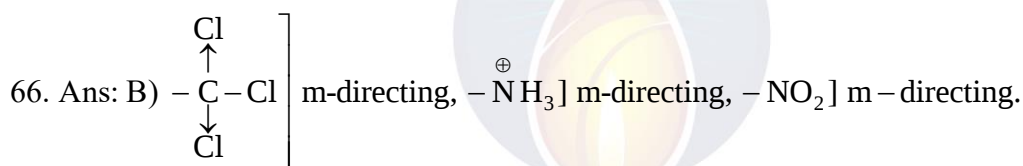
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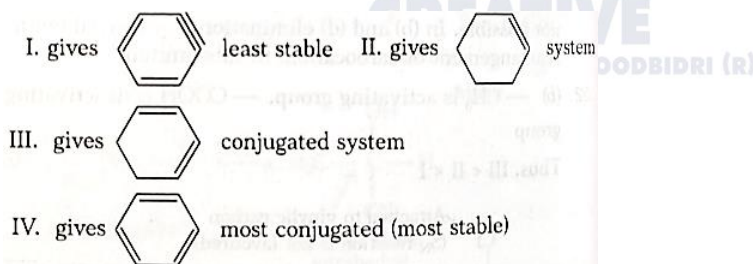
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SOLUTIONS

Reagent	Reaction with propene
A. $O_3 / Zn + H_2O$	$ \begin{array}{c} \text{CH}_3-\text{CH}=\text{CH}_2 \xrightarrow{\text{O}_3} \text{HCHO} \\ \text{O} \quad \text{O} \quad \text{O} \\ \text{O} \quad \text{O} \quad \text{O} \\ \text{O} \quad \text{O} \quad \text{O} \end{array} $ Formaldehyde + CH_3CHO Acetaldehyde
B. $\text{KMnO}_4 / \text{H}^+$	$ \text{CH}_3\text{CH}=\text{CH}_2 \xrightarrow[\text{H}^+]{\text{KMnO}_4} \text{CH}_3\text{COOH} + \text{CO}_2 $ Acetic acid
C. $\text{KMnO}_4 / \text{OH}^-$	$ \text{CH}_3-\text{CH}=\text{CH}_2 \xrightarrow[\text{OH}^-]{\text{KMnO}_4} \text{CH}_3-\text{CH}(\text{OH})-\text{CH}_2\text{OH} $ Propan-1, 2-diol
D. $\text{H}_2\text{O} / \text{H}^+$	$ \text{CH}_3\text{CH}=\text{CH}_2 \xrightarrow[\text{(Markownikoff's)}]{\text{H}_2\text{O}/\text{H}^+} \text{CH}_3-\text{CH}(\text{OH})-\text{CH}_3 $ Propan-2-ol
E. $\text{B}_2\text{H}_6 / \text{NaOH}$ and H_2O_2	$ \text{CH}_3\text{CH}=\text{CH}_2 \xrightarrow[\text{(ii) NaOH, H}_2\text{O}_2]{\text{(i) B}_2\text{H}_6} \text{CH}_3\text{CH}_2\text{CH}_2\text{OH} $ (Hydroboration oxidation) Propan-1-ol



67. Ans: A) Dehydration of



68. Ans: C) $\Delta T_b = i \times K_b \times M = 3 \times 0.52 \times \frac{13.44}{134.4} = 0.16$

69. Ans: D) Two solutions are isotonic if they have same molar concentration of particles.

In (D),

$0.1 \text{ M Ba}(\text{NO}_3)_2 = 0.3 \text{ M ion conc.}$

$0.1 \text{ M Na}_2\text{SO}_4 = 0.3 \text{ M ion conc.}$

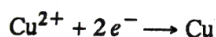
70. Ans: D) By definition of Kohlrausch law in terms of molar conductivities.



71. Ans: B

200 ml of 0.5 M CuSO_4 contains CuSO_4

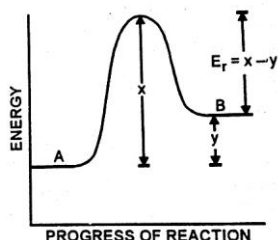
$$= 0.1 \text{ mole.}$$



To deposit 0.1 mole of Cu^{2+} , electricity required
 $= 2 \times 9650 \text{ C.}$

72. Ans: B

73. Ans: B) For an endothermic reaction.



74. Ans: A) N forms compounds from -3 to $+5$ oxidation states.

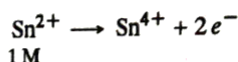
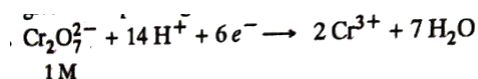
75. Ans: B) The increasing order of O – O bond length in O_3 , H_2O_2 and O_2 is O_2 (121 pm) < O_3 (128 pm) < H_2O_2 (148 pm)

In H_2O_2 , there is a single bond between O – O, in O_3 there is a resonance between a single and a double bond and in O_2 there is a double bond.

76. Ans: B) As the multiplicity of the bond increases, bond length decreases. Since N, O and F belong to the same period, their bond length follows the sequence $\text{N}_2 < \text{O}_2 < \text{F}_2$. Further, since Cl is bigger in size than F, therefore, bond length in F_2 is smaller than that in Cl_2 .

77. Ans: D) The correct statement is that transition elements exhibit variable valency.

78. Ans: A



1 M of Sn^{2+} in the process of oxidation gives $2e^-$

Now $6e^-$ reduce 1 M of $\text{Cr}_2\text{O}_7^{2-}$ completely.

Then $2e^-$ will reduce $\frac{M}{6} \times 2 = \frac{M}{3}$

i.e., $\frac{1}{3}$ moles of $\text{K}_2\text{Cr}_2\text{O}_7$.

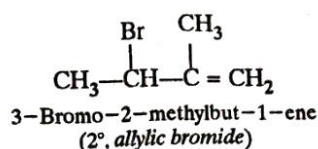
79. Ans: D) Pyridine is a monodentate ligand.

80. Ans: C) $[\text{Fe}(\text{CN})_6]^{3-}$ is paramagnetic but $[\text{Fe}(\text{CN})_6]^{4-}$ is diamagnetic though both are octahedral in shape.

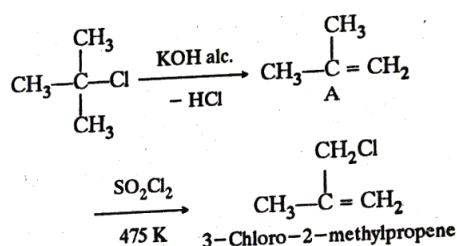


81. Ans: A) The formula of the complex will be $[\text{Co}(\text{NH}_3)_5\text{Cl}]\text{Cl}_2$.

82. Ans: A



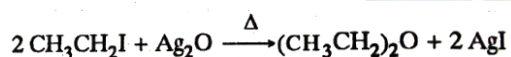
83. Ans: A



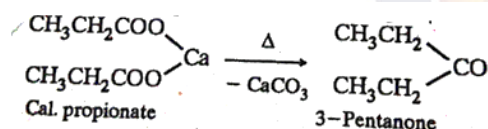
84. Ans: A

85. Ans: A

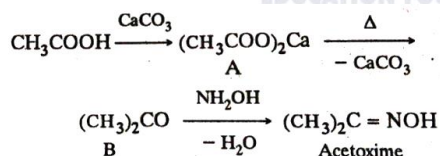
86. Ans: B



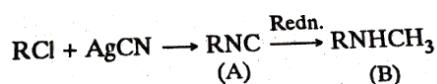
87. Ans: A



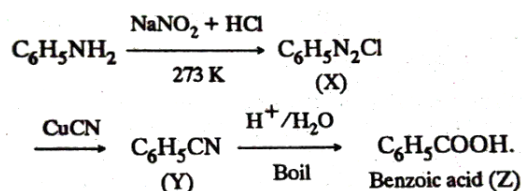
88. Ans: D



89. Ans: B



90. Ans: C





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Ans key	D	D	C	C	C	D	C	B	A	D
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Q. No	156	157	158	159	160	161	162	163	164	165
Ans key	C	B	A	C	A	C	A	C	A	C
Q. No	166	167	168	169	170	171	172	173	174	175
Ans key	A	B	A	D	B	C	D	C	C	D
Q. No	176	177	178	179	180					
Ans key	C	A	B	A	B					

